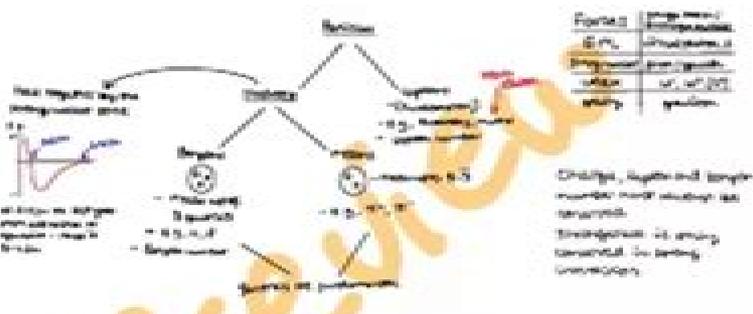


# Particles and Quantum



**Photoelectric effect:** When light strikes a metal surface, electrons are emitted. The energy of the photon must be greater than the work function of the metal.

**Equation:**  $E_{photon} = hf$

**Work function:**  $\phi$

**Maximum kinetic energy of electrons:**  $E_{kmax} = hf - \phi$

**Compton effect:** When a photon strikes a free electron, the photon's wavelength increases (energy decreases) and the electron recoils.

**Equation:**  $\lambda' - \lambda = \frac{h}{m_e c} (1 - \cos \theta)$

**Excitation:** An atom's electron jumps to a higher energy level by absorbing a photon.

**Alpha particles:** Helium nucleus ( $^4_2\text{He}$ ).

**Beta minus:** A neutron decays into a proton and an electron.

**Equation:**  $n \rightarrow p + e^- + \bar{\nu}_e$

**Beta plus:** A proton decays into a neutron and a positron.

**Equation:**  $p \rightarrow n + e^+ + \nu_e$

**Gamma:** High energy electromagnetic radiation.

**Equation:**  $E = hf$

**Pair production:** A photon with energy  $\geq 1.02 \text{ MeV}$  can create an electron-positron pair.

**Equation:**  $\gamma \rightarrow e^- + e^+$

**Annihilation:** An electron and a positron annihilate, producing two gamma photons.

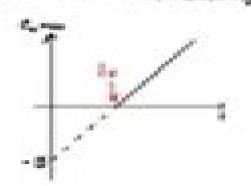
**Equation:**  $e^- + e^+ \rightarrow 2\gamma$



## Photoelectric effect

Use photons of sufficient energy on electrons on the surface of the metal, liberating them.

- 1. Photon absorbed** -  $hf$  (must be  $\geq \phi$ )
- 2. Emission** -  $hf - \phi$
- 3. Work function** - minimum energy to liberate electrons
- 4. Stopped potential** - minimum potential needed to stop electrons



The photoelectric effect provides the particle nature of light due to the discrete transfer of energy and momentum. It is not a wave phenomenon as it is not dependent on intensity of light.

**Stopping potential:** The potential needed to stop any electrons from reaching the other side.

**Equation:**  $eV_s = hf - \phi$

- 1. Change in momentum**
- 2. Stopping potential**

## De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of particles was proven by diffraction of electrons.

**Equation:**  $\lambda = \frac{h}{mv}$

**1. de Broglie wavelength**

**2. momentum / mass x velocity**

## Fluorescent tubes

These tubes contain mercury vapor. A high voltage is applied. This excites the mercury atoms, which then emit light as they return to the ground state.

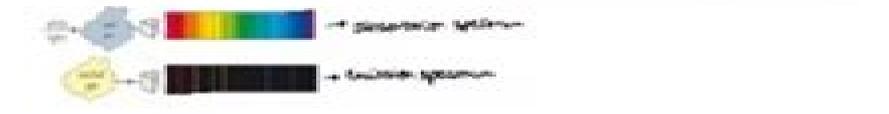
The gas contains phosphor which fluoresces in many colors. This is caused by higher energy levels when the electrons descend, say, from  $n=4$  to  $n=1$ .

A phosphor coating is on the inside of the tube and emits visible light.

## Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons or by collisions with other electrons.

**Ionization energy:** is enough energy to remove an electron, so will completely change the atom.



# A Level Physics Notes

**Huangqi Zhang**



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