

# Honors Chemistry Final Exam Study Guide

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**Mole - Answer-** SI unit for measuring the amount of substance;

**Molar Mass - Answer-** the mass of a mole of an element

**Representative particles - Answer-** refers to the species present in a substance: usually atoms, molecules, or formula units

**Avogadro's number - Answer-** the number of representative particles in a mole

**Avogadro's hypothesis - Answer-** states that equal volumes of gases at the same temperature and pressure contain equal numbers of particles

**Molar volume - Answer-** the quantity, 22.4 L

**Dimensional analysis - Answer-** mole to mass, mole to volume

**Percent Composition - Answer-** percent by mass of each element in the compound by which the relative amounts of the elements in a compound are expressed

**STP (Standard Temperature and Pressure) - Answer-** a temperature of 0°C and a pressure of 101.3 kPa, or 1 atmosphere (atm)

**Empirical formula - Answer-** the basic ratio gives the lowest whole-number ratio of the atoms of the elements contained in the compound

**Molecular formula - Answer-** either the same as its experimentally determined empirical formula, or it is a simple whole-number multiple of its empirical formula

**Types of reactions - Answer-** the five general types of reaction are combination, decomposition, single-replacement, double-replacement, and combustion

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