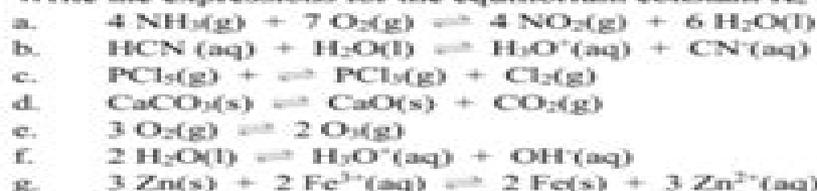


Chemical Equilibrium**STUDY QUESTIONS & PROBLEMS**

1. Write the expressions for the equilibrium constant
- K_c
- for the following reactions:



2. Write the equilibrium constant expressions for the following reactions. How are they related to one another?



3. Calculate the value of the equilibrium constant for the following system, given the data shown:



Concentrations at equilibrium:

$[\text{H}_2]$	=	1.5 mol Liter ⁻¹
$[\text{CO}_2]$	=	2.5 mol Liter ⁻¹
$[\text{H}_2\text{O}]$	=	0.5 mol Liter ⁻¹
$[\text{CO}]$	=	3.0 mol Liter ⁻¹

4. Chlorine molecules will dissociate at high temperatures into chlorine atoms. At 3000°C, for example, K_c for the equilibrium shown is 0.55. If the partial pressure of chlorine molecules is 1.5 atm, calculate the partial pressure of the chlorine atoms:

$$\text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{Cl}(\text{g})$$
5. Suppose that 0.50 moles of hydrogen gas, 0.50 moles of iodine gas, and 0.75 moles of hydrogen iodide gas are introduced into a 2.0 Liter vessel and the system is allowed to reach equilibrium.

$$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2 \text{HI}(\text{g})$$

 Calculate the concentrations of all three substances at equilibrium. At the temperature of the experiment, K_c equals 2.0×10^{-2} .
6. If the mechanism of a chemical equilibrium consists of two reversible elementary steps, each with its own equilibrium constant K_{c1} and K_{c2} , what expression relates the equilibrium constant K_c for the overall equilibrium to the two constants K_{c1} and K_{c2} ?
7. When 2.0 mol of carbon disulfide and 4.0 mol of chlorine are placed in a 1.0 Liter flask, the following equilibrium system results. At equilibrium, the flask is found to contain 0.30 mol of carbon tetrachloride. What quantities of the other components are present in this equilibrium mixture?



Chemical Equilibrium Practice Problems And Solutions

**Peter Atkins, Julio dePaula, Ron
Friedman**



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Chemical Engineering Principles and Applications Nuggenhalli S. Nandagopal, 2023-05-26 This text provides a clear and concise understanding of the principles and applications of chemical engineering using a rigorous yet easy to follow presentation The coverage is broad and it includes all the relevant concepts such as mass and energy balances mass transfer chemical reaction engineering and many more Elucidation of the principles is further reinforced by examples and practice problems with detailed solutions Firmly grounded in the fundamentals the book maximizes readers capacity to take on new problems and challenges in the field with confidence and conviction Providing a ready reference and review of essential principles and their applications in chemical engineering the book is ideal for undergraduate chemical engineering students as well as practicing engineers preparing for the engineering license exams FE and PE in USA and abroad **Student's Solutions Manual to Accompany Atkins' Physical Chemistry** C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 This solutions manual provides the authors detailed solutions to exercises and problems in physical chemistry It comprises solutions to exercises at the end of each chapter and solutions to numerical theoretical and additional problems **Partial Solutions**

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