

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Yicheng Fang



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Table of Contents Chemistry Unit 3 Energy Study Answers

1. Understanding the eBook Chemistry Unit 3 Energy Study Answers
 - The Rise of Digital Reading Chemistry Unit 3 Energy Study Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 3 Energy Study Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 3 Energy Study Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 3 Energy Study Answers
 - Personalized Recommendations

- Chemistry Unit 3 Energy Study Answers User Reviews and Ratings
- Chemistry Unit 3 Energy Study Answers and Bestseller Lists
- 5. Accessing Chemistry Unit 3 Energy Study Answers Free and Paid eBooks
 - Chemistry Unit 3 Energy Study Answers Public Domain eBooks
 - Chemistry Unit 3 Energy Study Answers eBook Subscription Services
 - Chemistry Unit 3 Energy Study Answers Budget-Friendly Options
- 6. Navigating Chemistry Unit 3 Energy Study Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 3 Energy Study Answers Compatibility with Devices
 - Chemistry Unit 3 Energy Study Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 3 Energy Study Answers
 - Highlighting and Note-Taking Chemistry Unit 3 Energy Study Answers
 - Interactive Elements Chemistry Unit 3 Energy Study Answers
- 8. Staying Engaged with Chemistry Unit 3 Energy Study Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 3 Energy Study Answers
- 9. Balancing eBooks and Physical Books Chemistry Unit 3 Energy Study Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 3 Energy Study Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 3 Energy Study Answers
 - Setting Reading Goals Chemistry Unit 3 Energy Study Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 3 Energy Study Answers
 - Fact-Checking eBook Content of Chemistry Unit 3 Energy Study Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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