

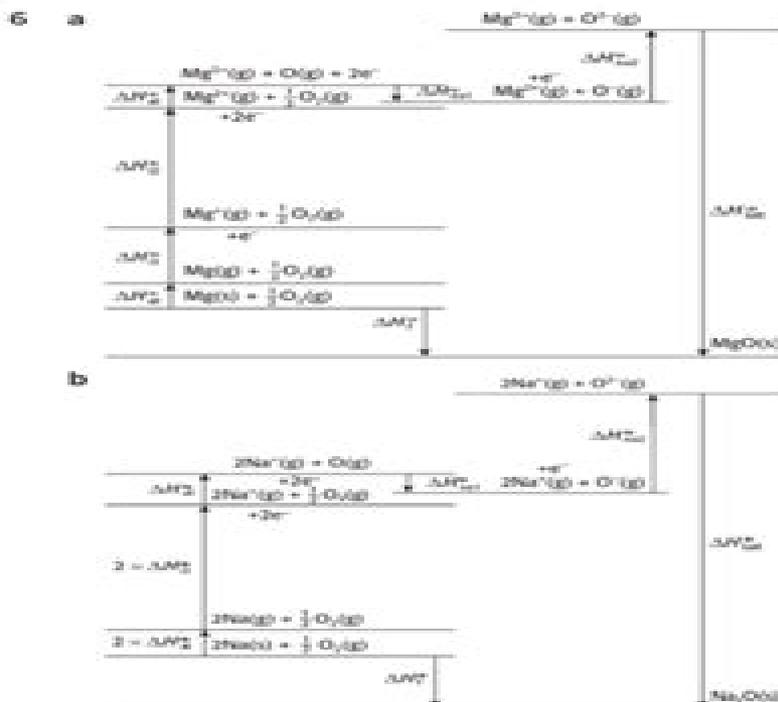
# Answers to SAQs

## Chapter 19

- 1 a standard temperature = 298 K, standard pressure =  $1 \times 10^5$  Pa / 100 kPa
- b i  $\text{Mg}^{2+}(\text{g}) + \text{O}^{2-}(\text{g}) \rightarrow \text{MgO}(\text{s})$   
 ii  $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$   
 iii  $2\text{Na}^+(\text{g}) + \text{S}^{2-}(\text{g}) \rightarrow \text{Na}_2\text{S}(\text{s})$
- 2 a The bond energy for chlorine is the enthalpy change  $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$ . The enthalpy change of atomisation is  $\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{Cl}(\text{g})$ . So the enthalpy change of atomisation is  $+244/2 = +122 \text{ kJ mol}^{-1}$ .
- b i  $\frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{O}(\text{g})$   
 ii  $\text{Ba}(\text{s}) \rightarrow \text{Ba}(\text{g})$   
 iii  $\frac{1}{2}\text{Br}_2(\text{l}) \rightarrow \text{Br}(\text{g})$
- c  $0 \text{ kJ mol}^{-1}$   
 Because helium exists naturally as single gaseous atoms, no change is involved in the process  $\text{He}(\text{g}) \rightarrow \text{He}(\text{g})$ .
- 3 a There must be an input of energy to overcome the repulsive forces between the (negative) electron and the negative ion.
- b  $+440 \text{ kJ mol}^{-1}$
- c i  $\text{I}(\text{g}) + \text{e}^- \rightarrow \text{I}^-(\text{g})$   
 ii  $\text{S}^-(\text{g}) + \text{e}^- \rightarrow \text{S}^{2-}(\text{g})$
- 4 a i  $\text{Cs}(\text{g}) \rightarrow \text{Cs}^+(\text{g}) + \text{e}^-$   
 ii  $\text{Al}^{3+}(\text{g}) \rightarrow \text{Al}^{2+}(\text{g}) + \text{e}^-$   
 iii  $\text{Ca}(\text{s}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CaO}(\text{s})$   
 iv  $\text{Fe}(\text{s}) + 1\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{FeCl}_3(\text{s})$
- b  $\Delta H_{\text{form}}^{\ominus} = \Delta H_{\text{at}}^{\ominus} - [\Delta H_{\text{ion}}^{\ominus}(\text{Na}) + \Delta H_{\text{ion}}^{\ominus}(\text{Cl}) + \Delta H_{\text{at}}^{\ominus}(\frac{1}{2}\text{Cl}_2(\text{g})) + \Delta H_{\text{form}}^{\ominus}(\text{Cl})]$   
 $\Delta H_{\text{form}}^{\ominus} = (-411) - [(+107) + (+496) + (+122) + (-348)]$   
 $= +377 \text{ kJ mol}^{-1}$   
 $\Delta H_{\text{ion}}^{\ominus} = (-411) - (+377) = -788 \text{ kJ mol}^{-1}$



- b i enthalpy change of atomisation of iodine  
 ii first electron affinity of nitrogen  
 iii enthalpy change of formation of strontium chloride  
 iv lattice energy of cadmium chloride



# Chemistry Chapter 19 Assessment Answers

**Richard H. Langley, John Moore**



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an essay that seemed the apex of what I was capable of writing. Although the topic of that essay is now lost to my memory, the grade I received was not a B. The best I could do was a B. The realization sank in that maybe I was not such a good writer. In those days, my youthful hubris did not understand about the capacity of being. In other words, being challenged would result in my intellectual growth, an academic restatement of Nietzsche's "What does not destroy me makes me stronger." Consequently, I asked to be withdrawn from Gifted and Talented English in the eighth grade.

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*Chemistry* Stanley R. Radel, Marjorie H. Navidi, 1990

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