

Chapter 10 test**Matching**

Match each item with the correct statement below.

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

- _____ 1. the number of representative particles of a substance present in 1 mole of that substance
- _____ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- _____ 3. the SI unit used to measure amount of substance
- _____ 4. 0°C and 1 atm
- _____ 5. the percent by mass of each element in a compound
- _____ 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 7. What SI unit is used to measure the number of representative particles in a substance?
- | | |
|-------------|-----------|
| a. kilogram | c. kelvin |
| b. ampere | d. mole |
- _____ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C₃H₇O?
- | | |
|-------------------------------------|--------------------------------------|
| a. $5 \times (6.02 \times 10^{23})$ | c. 35 |
| b. 5 | d. $35 \times (6.02 \times 10^{23})$ |
- _____ 9. All of the following are equal to Avogadro's number EXCEPT _____
- | |
|--|
| a. the number of atoms of bromine in 1 mol Br ₂ |
| b. the number of atoms of gold in 1 mol Au |
| c. the number of molecules of nitrogen in 1 mol N ₂ |
| d. the number of molecules of carbon monoxide in 1 mol CO |
- _____ 10. How many moles of tungsten atoms are in 4.8×10^{23} atoms of tungsten?
- | | |
|-------------------------------|-------------------------------|
| a. 8.0×10^{-2} moles | c. 1.3×10^{-1} moles |
| b. 8.0×10^1 moles | d. 1.3×10^{-2} moles |
- _____ 11. How many atoms are in 0.075 mol of titanium?
- | | |
|-------------------------|-------------------------|
| a. 1.2×10^{23} | c. 6.4×10^2 |
| b. 2.2×10^{24} | d. 4.5×10^{22} |
- _____ 12. How many molecules are in 2.10 mol CO₂?
- | | |
|------------------------------------|-------------------------------------|
| a. 2.53×10^{24} molecules | c. 3.49×10^{-24} molecules |
| b. 3.79×10^{24} molecules | d. 1.26×10^{24} molecules |

Chemistry Chapter 10 Test Answers

M Walker



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web these structures also known as lewis structures or electron dot structures are drawings that visually demonstrate how electrons are shared and arranged around atoms the electrons denoted as dots are called lone pairs and belong to an individual atom electrons denoted as lines are bonds and show the sharing of two electrons between two

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