

## Chapter 9 Practice Problems: Chemical Names and Formulas

### Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
  - oxygen
  - iodine
  - sodium
  - aluminum
  - nickel, two electrons lost
  - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
  - $\text{Cr}^{3+}$
  - $\text{P}^{3-}$
  - $\text{Li}^+$
  - $\text{Ca}^{2+}$
  - $\text{Cl}^-$
  - $\text{O}^{2-}$
- Name each ion. Identify each as a cation or an anion.
  - $\text{Sn}^{2+}$
  - $\text{Co}^{3+}$
  - $\text{Br}^-$
  - $\text{K}^+$
  - $\text{H}^-$
  - $\text{Mn}^{2+}$
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
  - carbonate ion
  - nitrite ion
  - sulfate ion
  - hydroxide ion
  - chromate ion
  - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
  - $\text{CN}^-$
  - $\text{HCO}_3^-$
  - $\text{PO}_4^{3-}$
  - $\text{Cl}^-$
  - $\text{Ca}^{2+}$
  - $\text{SO}_3^{2-}$

### 9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
  - magnesium oxide
  - tin(II) fluoride
  - potassium iodide
  - aluminum chloride
  - sodium sulfide
  - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
  - $\text{Ba}^{2+}$  and  $\text{Cl}^-$
  - $\text{Ag}^+$  and  $\text{I}^-$
  - $\text{Ca}^{2+}$  and  $\text{S}^{2-}$
  - $\text{K}^+$  and  $\text{Br}^-$
  - $\text{Al}^{3+}$  and  $\text{O}^{2-}$
  - $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
- Name the following binary ionic compounds.
  - $\text{MnO}_2$
  - $\text{Li}_3\text{N}$
  - $\text{CaCl}_2$
  - $\text{SrBr}_2$
  - $\text{NiCl}_2$
  - $\text{K}_2\text{S}$
  - $\text{Fe}_3\text{N}_2$
  - $\text{SnCl}_4$
- Write formulas for the following ionic compounds.
  - sodium phosphate
  - magnesium sulfate
  - sodium hydroxide
  - potassium cyanide
  - ammonium chloride
  - potassium dichromate
- Write formulas for the following ionic compounds.
  - $\text{NH}_4^+$  and  $\text{SO}_4^{2-}$
  - $\text{K}^+$  and  $\text{NO}_3^-$
  - barium ion and hydroxide ion
  - lithium ion and carbonate ion

# Chapter 9 Chemical Names Formulas Practice Problems Answers

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**Chapter 9 Chemical Names Formulas Practice Problems Answers:**

## **Chapter 9 Chemical Names Formulas Practice Problems Answers** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become much more evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **Chapter 9 Chemical Names Formulas Practice Problems Answers**, a literary masterpiece that delves deep to the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall affect readers.

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### **Table of Contents Chapter 9 Chemical Names Formulas Practice Problems Answers**

1. Understanding the eBook Chapter 9 Chemical Names Formulas Practice Problems Answers
  - The Rise of Digital Reading Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 9 Chemical Names Formulas Practice Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Personalized Recommendations
  - Chapter 9 Chemical Names Formulas Practice Problems Answers User Reviews and Ratings
  - Chapter 9 Chemical Names Formulas Practice Problems Answers and Bestseller Lists

5. Accessing Chapter 9 Chemical Names Formulas Practice Problems Answers Free and Paid eBooks
  - Chapter 9 Chemical Names Formulas Practice Problems Answers Public Domain eBooks
  - Chapter 9 Chemical Names Formulas Practice Problems Answers eBook Subscription Services
  - Chapter 9 Chemical Names Formulas Practice Problems Answers Budget-Friendly Options
6. Navigating Chapter 9 Chemical Names Formulas Practice Problems Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 9 Chemical Names Formulas Practice Problems Answers Compatibility with Devices
  - Chapter 9 Chemical Names Formulas Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Highlighting and Note-Taking Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Interactive Elements Chapter 9 Chemical Names Formulas Practice Problems Answers
8. Staying Engaged with Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 9 Chemical Names Formulas Practice Problems Answers
9. Balancing eBooks and Physical Books Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 9 Chemical Names Formulas Practice Problems Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Setting Reading Goals Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Fact-Checking eBook Content of Chapter 9 Chemical Names Formulas Practice Problems Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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