



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{CO}_2$
  - $\text{H}_2\text{S}$
  - $\text{CaCl}_2$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{SO}_3$
  - $\text{H}_2\text{SO}_4$

- Classify each of the following as either ionic or covalent compound. Write the name of the compound.
- Write the formula for each of the following ionic compounds. Write the name of the compound.
- Write the formula for each of the following covalent compounds. Write the name of the compound.

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for methane.  $\text{C}$
- Draw the Lewis structure for carbon dioxide.  $\text{C}$
- Draw the Lewis structure for acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for hydrogen cyanide.  $\text{H}-\text{C}\equiv\text{N}$



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital do you combine to form the sigma bond in the hydrogen molecule  $\text{H}_2$ ?
- Which orbital and sub-orbital do you combine to form the sigma bond in  $\text{H}_2$ ?

# Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers

**G Thomas**



## Chapter 8 Covalent Bonding Section 81 Molecular Compounds Answers:

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