

## Chapter 8: Covalent Bonding

Students should:

- 1.14 know what is meant by the term molecule
- 1.44 know that a covalent bond is formed between atoms by the sharing of a pair of electrons
- 1.45 understand covalent bonds in terms of electrostatic attractions
- 1.46 understand how to use dot-and-cross diagrams to represent covalent bonds in:
  - diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides
  - inorganic molecules including water, ammonia and carbon dioxide
  - organic molecules containing up to two carbon atoms, including methane, ethane, ethene and those containing halogen atoms
- 1.47 explain why substances with simple molecular structures are gases or liquids, or solids with low melting and boiling points  
*the term intermolecular forces of attraction can be used to represent all forces between molecules*
- 1.48 explain why the melting and boiling points of substances with simple molecular structures increase, in general, with increasing relative molecular mass
- 1.49 explain why substances with giant covalent structures are solids with high melting and boiling points
- 1.50 explain how the structures of diamond, graphite and C<sub>60</sub> fullerene influence their physical properties, including electrical conductivity and hardness
- 1.51 know that covalent compounds do not usually conduct electricity

### 8.1 Covalent bonding

A covalent bond is resulted when non-metal atoms combine with non-metal atoms by sharing of outermost shell electrons. Strong, directional electrostatic attraction is formed between the (negatively charged) shared electrons and the (positively charged) nuclei of the two bonded atoms. Such attraction is called a



The bonded atoms can then attain an octet or duplet of electrons and become stable (i.e. each atom attains the stable electronic structure of its nearest noble gas).

# Chapter 8 Covalent Bonding Answers Key Dongmengore

**JL Elias**



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## **Table of Contents Chapter 8 Covalent Bonding Answers Key Dongmengore**

1. Understanding the eBook Chapter 8 Covalent Bonding Answers Key Dongmengore
  - The Rise of Digital Reading Chapter 8 Covalent Bonding Answers Key Dongmengore
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Answers Key Dongmengore
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Covalent Bonding Answers Key Dongmengore
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Answers Key Dongmengore
  - Personalized Recommendations
  - Chapter 8 Covalent Bonding Answers Key Dongmengore User Reviews and Ratings
  - Chapter 8 Covalent Bonding Answers Key Dongmengore and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Answers Key Dongmengore Free and Paid eBooks
  - Chapter 8 Covalent Bonding Answers Key Dongmengore Public Domain eBooks
  - Chapter 8 Covalent Bonding Answers Key Dongmengore eBook Subscription Services
  - Chapter 8 Covalent Bonding Answers Key Dongmengore Budget-Friendly Options
6. Navigating Chapter 8 Covalent Bonding Answers Key Dongmengore eBook Formats

- ePub, PDF, MOBI, and More
  - Chapter 8 Covalent Bonding Answers Key Dongmengore Compatibility with Devices
  - Chapter 8 Covalent Bonding Answers Key Dongmengore Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Highlighting and Note-Taking Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Interactive Elements Chapter 8 Covalent Bonding Answers Key Dongmengore
  8. Staying Engaged with Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 8 Covalent Bonding Answers Key Dongmengore
  9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Answers Key Dongmengore
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Setting Reading Goals Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Answers Key Dongmengore
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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