

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



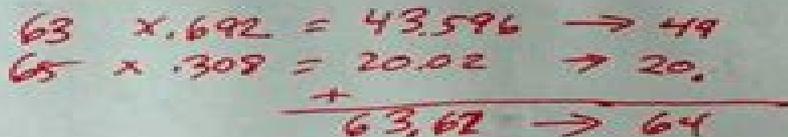
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



Chapter 6 Assessment Chemistry Answers

John T. Moore, Richard H. Langley



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Table of Contents Chapter 6 Assessment Chemistry Answers

1. Understanding the eBook Chapter 6 Assessment Chemistry Answers
 - The Rise of Digital Reading Chapter 6 Assessment Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Assessment Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Assessment Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Assessment Chemistry Answers

- Personalized Recommendations
- Chapter 6 Assessment Chemistry Answers User Reviews and Ratings
- Chapter 6 Assessment Chemistry Answers and Bestseller Lists
- 5. Accessing Chapter 6 Assessment Chemistry Answers Free and Paid eBooks
 - Chapter 6 Assessment Chemistry Answers Public Domain eBooks
 - Chapter 6 Assessment Chemistry Answers eBook Subscription Services
 - Chapter 6 Assessment Chemistry Answers Budget-Friendly Options
- 6. Navigating Chapter 6 Assessment Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Assessment Chemistry Answers Compatibility with Devices
 - Chapter 6 Assessment Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Assessment Chemistry Answers
 - Highlighting and Note-Taking Chapter 6 Assessment Chemistry Answers
 - Interactive Elements Chapter 6 Assessment Chemistry Answers
- 8. Staying Engaged with Chapter 6 Assessment Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Assessment Chemistry Answers
- 9. Balancing eBooks and Physical Books Chapter 6 Assessment Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Assessment Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 6 Assessment Chemistry Answers
 - Setting Reading Goals Chapter 6 Assessment Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 6 Assessment Chemistry Answers

- Fact-Checking eBook Content of Chapter 6 Assessment Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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