

Chapter 11 Nuclear Chemistry

Review Questions

Classical Energy is energy stored in the bonds of atoms and molecules. **Nuclear Energy** is stored throughout the nucleus of atoms changing them into different atoms. Nuclear energy is produced by **radioactivity**.

The Nucleus



Nucleus is composed of two main particles, **protons** and **neutrons**.
The number of protons is the **atomic number**.
The number of protons and neutrons together is the **mass number**.

Isotopes that have the same number of protons but different mass, but identical chemical properties are called **isotopes**.

Radioactivity is the emission of an element's characteristic particles, which can be called **alpha rays**, **beta rays**, or **gamma rays**.

Table 11.1: Properties of alpha, beta and gamma rays		
	alpha rays	beta rays
Mass	4	1/1836
Charge	2+	1-
Penetration power	Low (10 ² cm)	High (10 ⁴ cm)
Ionisation power	High	Low
Deflection	Deflected	Deflected

Types of Radioactive Rays

- Alpha rays** is the least ionising particles



- Beta rays** is the least ionising particles in highly ionised elements, produced by the nucleus: ${}^0_{-1}\text{e}$



- Gamma rays** is the most ionising, which is in highly ionising substances that release energy corresponding to the loss of a nucleon particle: γ

Chapter 21 Nuclear Chemistry Mixed Review Answers

AW Chickering



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