

Chapter 6 Thermochemistry

Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ.
 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is 0.71 J/°C·g. Calculate the energy required to:
 - Raise the temperature of 1.0 mol graphite by 1.0 °C.
 - Raise the temperature of 850 g graphite by 150 °C.
 - Raise the temperature of 75 kg graphite from 294 K to 348 K.
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at 24.6 °C. After the reaction, the final temperature is 31.3 °C. Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of 4.18 J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($C_{15}H_{26}O$) has an energy of combustion of -5903.6 kJ/mol. A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by 2.28 °C. Calculate the heat capacity of the calorimeter.
A 0.1964 g sample of quinone ($C_{10}H_6O_2$) is burned in a bomb calorimeter that has a heat capacity of 1.56 kJ/°C. The temperature of the calorimeter increases by 3.2 °C. Calculate the energy of combustion of quinone per gram and per mole.

Chapter 6 Thermochemistry Homework Problems

Steven S. Zumdahl, Kenneth C. Brooks



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and equipment Practice drills at the end of each content review chapter plus step by step walk throughs of sample exam questions

Solutions Manual to Accompany General Chemistry with Qualitative Analysis, Second Edition Ralph H. Petrucci, Robert K. Wismer, 1987

Chemistry Steven S. Zumdahl, Kenneth C. Brooks, 1986

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16

Chemistry The Molecular Nature of Matter and Change CHEMISTRY SILBERBERG, 2003

Rocket Propulsion Stephen D. Heister, William E. Anderson, Timothée L. Pourpoint, R. Joseph Cassady, 2019-02-07

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Schaum's Outline of Theory and Problems of College Chemistry Jerome Laib Rosenberg, Lawrence M. Epstein, 1990

Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993

Introduction to Supercritical Fluids Richard Smith, Hiroshi Inomata, Cor Peters, 2013-12-08

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