

Name: \_\_\_\_\_ Pd: \_\_\_\_\_

## Lab - Evidence for Chemical Change

One way of knowing that a chemical change has occurred is to observe that the properties of the products are different from those of the reactants. The new product can then become a reactant in another chemical reaction. In this experiment you will observe a sequence of changes that occur when a solution that begins as copper (II) nitrate is treated with a series of different reactants. All of the reactions will take place in the same test tube. At each step you will look for evidence that a new substance is formed as a result of a chemical change. You will also observe heat changes and relate them to chemical reactions.

### Pre-Lab

For each reaction: **IDENTIFY** the type of chemical reaction, **PREDICT** the products, and **BALANCE** the chemical equation.



**Type of Reaction:**



**Type of Reaction:**



**Type of Reaction:**



**Type of Reaction:**



**Type of Reaction:**

### Procedure

- Place 50 mL of water into the 100 mL beaker and heat it on a hot plate until boiling. This will be the hot-water bath you will use in **step 6**.
- While the water bath is heating, use the lab marker and ruler to make three marks on the test tube that are 1 cm apart. Make the marks starting at the bottom of the test tube and moving toward the top.
- Add 1.0 M  $\text{Cu}(\text{NO}_3)_2$ , copper (II) nitrate, to the **first** mark on the test tube as in **Figure 1**.
- Add 1.0 M NaOH, sodium hydroxide, up to the **second** mark on the test tube as in **Figure 2**.

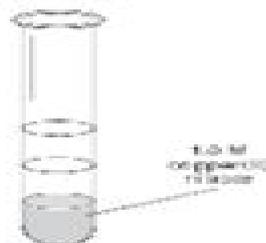


Figure 1

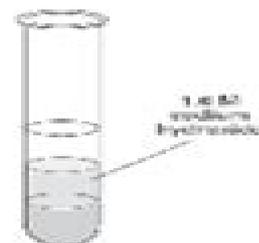


Figure 2

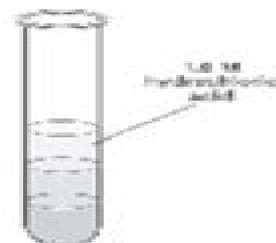


Figure 3

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