

Temperature

°C, K

measure of average KE  
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from  
an object of high E → low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or  $c_p$ ):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C.

change in heat = specific heat x mass x change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

$\Delta T$  in °C or K; m in g; c in J/g°C

Molar heat capacity = heat capacity in J/mol°C instead of J/g°C

Don't forget: the size of 1°C = 1 K, so K can be used in this calculation in place of °C.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C? ( $c_p$  of H<sub>2</sub>O = 4.18 J/g°C)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C? ( $c_p$  of gold = 0.129 J/g°C)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

# Chapter 6 Thermochemistry Test

**Robert K. Wismer**



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## **Table of Contents Chapter 6 Thermochemistry Test**

1. Understanding the eBook Chapter 6 Thermochemistry Test
  - The Rise of Digital Reading Chapter 6 Thermochemistry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Thermochemistry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Thermochemistry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Thermochemistry Test
  - Personalized Recommendations
  - Chapter 6 Thermochemistry Test User Reviews and Ratings
  - Chapter 6 Thermochemistry Test and Bestseller Lists
5. Accessing Chapter 6 Thermochemistry Test Free and Paid eBooks
  - Chapter 6 Thermochemistry Test Public Domain eBooks
  - Chapter 6 Thermochemistry Test eBook Subscription Services
  - Chapter 6 Thermochemistry Test Budget-Friendly Options
6. Navigating Chapter 6 Thermochemistry Test eBook Formats

- ePub, PDF, MOBI, and More
  - Chapter 6 Thermochemistry Test Compatibility with Devices
  - Chapter 6 Thermochemistry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 6 Thermochemistry Test
    - Highlighting and Note-Taking Chapter 6 Thermochemistry Test
    - Interactive Elements Chapter 6 Thermochemistry Test
  8. Staying Engaged with Chapter 6 Thermochemistry Test
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 6 Thermochemistry Test
  9. Balancing eBooks and Physical Books Chapter 6 Thermochemistry Test
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 6 Thermochemistry Test
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 6 Thermochemistry Test
    - Setting Reading Goals Chapter 6 Thermochemistry Test
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 6 Thermochemistry Test
    - Fact-Checking eBook Content of Chapter 6 Thermochemistry Test
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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