

16

SOLUTIONS

SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)

This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves.

► Solution Formation (pages 471–472)

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

- Underline the condition that causes sugar to dissolve *faster* in water.
 - as a whole cube or in granulated form?
 - when allowed to stand or when stirred?
 - at a higher temperature or a lower temperature?
- Name three factors that influence the rate at which a solute dissolves in a solvent.
 - agitation (stirring or shaking)
 - temperature
 - particle size
- Is the following sentence true or false? Finely ground particles dissolve more rapidly than larger particles because finer particles expose a greater surface area to the colliding solvent molecules. true

► Solubility (pages 472–473)

- Complete the following table showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

Procedure	Amount Dissolved	Amount Not Dissolved
Add 36.0 g of sodium chloride to the water	36.0 g	0.0 g
Add an additional 1.0 g of sodium chloride	0.2 g	0.8 g
Determine the total amount that dissolves	36.2 g	

- The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's solubility at that temperature.
- If a solution contains the *maximum* amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) saturated solution.

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Rachel Sandford



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