

# CHEM 103 Lab Exam 3 (100 out of 100) Questions and Verified Answers (Portage Learning)

## Scientific Calculator

### Question 1

5 / 10 pts

Answer the following two questions:

1. Why was the hydrate heated with the Fisher burner in part 2 of lab 3?
2. For what purpose was SEM carried out in lab 3?

1. The hydrate was heated with the Fisher burner in lab 3 to cause loss of water.
2. The SEM was carried out in lab 3 to detect the elemental composition.

2. The SEM was carried out in lab 3 to detect the elemental composition. (We did not use it for imaging)

### Question 2

10 / 10 pts

## Scientific Calculator

Answer the following question.

A 1.1000 gram hydrate sample chosen from  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ ,  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$  and  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  was heated and found to lose 0.6920 gram of  $\text{H}_2\text{O}$ . (1) Show the calculation of the %  $\text{H}_2\text{O}$  in the unknown hydrate sample. (2) Show the calculation of the %  $\text{H}_2\text{O}$  in each of the hydrate compounds and identify the unknown hydrate from the list.

Atomic weights: H = 1.008, O = 16.00. MWs:  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$  = 286.15,  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$  = 241.43,  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$  = 203.301 and  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  = 244.462

%  $\text{H}_2\text{O}$  in unknown =  $(0.6920 / 1.1000) \times 100 = 62.91\%$

%  $\text{H}_2\text{O}$  in  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$  =  $180.16 / 286.15 \times 100 = 62.96\%$

unknown is  **$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$**

### Question 3

10 / 10 pts

# Chem 110 Lab Questions And Answers

**Tobias Bleicker**



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