

1-6, 8, 9, 11

## Hydrolysis of Salts and pH of Buffer Solutions

## Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define Brønsted-Lowry acids and bases.

According to Brønsted-Lowry definitions, an acid is a proton donor and a base is a proton acceptor.

2

- Which of the following ions will react with water in a hydrolysis reaction:

$K^+$ ,  $Ba^{2+}$ ,  $Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$ ,  $SO_3^{2-}$ , and  $Cl^-$ ?

$Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$  and  $SO_3^{2-}$  will undergo hydrolysis.

3

- For those ions in question 2 that undergo hydrolysis, write net ionic equations for the hydrolysis reaction.

Example:  $Cu^{2+}(aq) + 2H_2O(l) \rightleftharpoons Cu(OH)_2(s) + 2H^+(aq)$ .

4

- The  $K_a$  for HCN is  $4.9 \times 10^{-10}$ . What is the value of  $K_b$  for  $CN^-$ ?

$$K_b = \frac{1.0 \times 10^{-14}}{4.9 \times 10^{-10}} = 2.0 \times 10^{-5}$$

5. What are the conjugate base and conjugate acid of  $H_2PO_4^-$ ?

The conjugate base is  $HPO_4^{2-}$  and the conjugate acid is  $H_3PO_4$ .

6

- From what acid and what base were the following salts made:  $K_2SO_4$ ,  $NH_4Cl$ , and  $Ca(NO_3)_2$ ?

For  $Ca^{2+}$  and  $SO_4^{2-}$  salt is made from  $Ca(OH)_2$  and  $H_2SO_4$ .

7

- Define the term *salt*.

# Ph Properties Of Buffer Solutions Answer Key Pre Lab

**Tobias Bleicker**



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## **Table of Contents Ph Properties Of Buffer Solutions Answer Key Pre Lab**

1. Understanding the eBook Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - The Rise of Digital Reading Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Advantages of eBooks Over Traditional Books
2. Identifying Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Personalized Recommendations
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab User Reviews and Ratings

- Ph Properties Of Buffer Solutions Answer Key Pre Lab and Bestseller Lists
- 5. Accessing Ph Properties Of Buffer Solutions Answer Key Pre Lab Free and Paid eBooks
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab Public Domain eBooks
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab eBook Subscription Services
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab Budget-Friendly Options
- 6. Navigating Ph Properties Of Buffer Solutions Answer Key Pre Lab eBook Formats
  - ePub, PDF, MOBI, and More
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab Compatibility with Devices
  - Ph Properties Of Buffer Solutions Answer Key Pre Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Highlighting and Note-Taking Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Interactive Elements Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 8. Staying Engaged with Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 9. Balancing eBooks and Physical Books Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Setting Reading Goals Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Fact-Checking eBook Content of Ph Properties Of Buffer Solutions Answer Key Pre Lab
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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