

SOLUTIONS
THERMODYNAMICS PRACTICE PROBLEMS FOR NON-TECHNICAL MAJORS

Thermodynamic Properties

1. If an object has a weight of 10 lbf on the moon, what would the same object weigh on Jupiter?

$$g_{\text{Jupiter}} = 75 \frac{\text{ft}}{\text{sec}^2} \quad g_{\text{Moon}} = 5.4 \frac{\text{ft}}{\text{sec}^2} \quad g_c = 32 \frac{\text{lbm-ft}}{\text{lbf-sec}^2}$$

$$|W| = \frac{mg}{g_c} \Rightarrow m = \frac{|W|_{\text{Moon}} g_c}{g_{\text{Moon}}} = \frac{10 \times 32}{5.4} = 59.26 \text{ lbm}$$

$$|W|_{\text{Jupiter}} = \frac{mg_{\text{Jupiter}}}{g_c} = \frac{59.26 \times 75}{32} = \boxed{139 \text{ lbf}}$$

2. An object that weighs 50 lbf on earth is moved to Saturn where its new weight is 105 lbf. What is the acceleration due to gravity on Saturn?

$$g_{\text{Earth}} = 32 \frac{\text{ft}}{\text{sec}^2} \quad g_c = 32 \frac{\text{lbm-ft}}{\text{lbf-sec}^2}$$

$$50 \text{ lbf on earth} \Rightarrow 50 \text{ lbm}$$

$$|W| = \frac{mg}{g_c} \Rightarrow g = \frac{|W| g_c}{m} = \frac{105 \times 32}{50} = \boxed{67.2 \frac{\text{ft}}{\text{sec}^2}}$$

3. Define, using equations, specific volume (v) and density (ρ). What is the mathematical relationship between these two terms?

$$v = \frac{V}{m}, \quad \rho = \frac{m}{V}, \quad v = \frac{1}{\rho} \text{ or } \rho = \frac{1}{v}$$

Temperature and Pressure Measurements

4. (a) Define temperature.
(b) What is the absolute temperature scale corresponding to Fahrenheit?
(c) Convert 100° F to that absolute scale.

(a) Temperature: a measure of molecular activity of a substance.

(b) Rankine

(c) °R = °F + 460 \Rightarrow 100° F converts to 560° R

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Lectures in Classical Thermodynamics with an Introduction to Statistical Mechanics Daniel Blankschtein, 2021-03-15 This textbook facilitates students ability to apply fundamental principles and concepts in classical thermodynamics to solve challenging problems relevant to industry and everyday life It also introduces the reader to the fundamentals of statistical mechanics including understanding how the microscopic properties of atoms and molecules and their associated intermolecular interactions can be accounted for to calculate various average properties of macroscopic systems The author emphasizes application of the fundamental principles outlined above to the calculation of a variety of thermodynamic properties to the estimation of conversion efficiencies for work production by heat interactions and to the solution of practical thermodynamic problems related to the behavior of non ideal pure fluids and fluid mixtures including phase equilibria and chemical reaction equilibria The book contains detailed solutions to many challenging sample problems in classical thermodynamics and statistical mechanics that will help the reader crystallize the material taught Class tested and perfected over 30 years of use by nine time Best Teaching Award recipient Professor Daniel Blankschtein of the Department of Chemical Engineering at MIT the book is ideal for students of Chemical and Mechanical Engineering Chemistry and Materials Science who will benefit greatly from in depth discussions and pedagogical explanations of key concepts Distills critical concepts methods and applications from leading full length textbooks along with the author s own deep understanding of the material taught into a concise yet rigorous graduate and advanced undergraduate text Enriches the standard curriculum with succinct problem based learning strategies derived from the content of 50 lectures given over the years in the Department of Chemical Engineering at MIT Reinforces concepts covered with detailed solutions to illuminating and challenging homework problems

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Fundamentals of Engineering Merle C. Potter,1999 **Introductory Chemical Engineering Thermodynamics** J. Richard Elliott,Carl T. Lira,2012-02-06 A Practical Up to Date Introduction to Applied Thermodynamics Including Coverage of Process Simulation Models and an Introduction to Biological Systems Introductory Chemical Engineering Thermodynamics Second Edition helps readers master the fundamentals of applied thermodynamics as practiced today with extensive development of molecular perspectives that enables adaptation to fields including biological systems environmental applications and nanotechnology This text is distinctive in making molecular perspectives accessible at the introductory level and connecting properties with practical implications Features of the second edition include Hierarchical instruction with increasing levels of detail Content requiring deeper levels of theory is clearly delineated in separate sections and chapters Early introduction to the overall perspective of composite systems like distillation columns reactive processes and biological systems Learning objectives problem solving strategies for energy balances and phase equilibria chapter summaries and important equations for every chapter Extensive practical examples especially coverage of non ideal mixtures which include water contamination via hydrocarbons polymer blending recycling oxygenated fuels hydrogen bonding osmotic pressure electrolyte solutions zwitterions and biological molecules and other contemporary issues Supporting software in formats for both MATLAB and spreadsheets Online supplemental sections and resources including instructor slides ConcepTests coursecast videos and other useful resources Thermodynamics Problem Solver The Editors of REA,Ralph Pike,2013-01-01 REA s Thermodynamics Problem Solver Each Problem Solver is an insightful and essential study and solution guide chock full of clear concise problem solving gems Answers to all of your questions can be found in one convenient source from one of the most trusted names in reference solution guides More useful more practical and more informative these study aids are the best review books and textbook companions available They re perfect for undergraduate and graduate studies This highly useful reference provides thorough coverage of pressure work and heat energy entropy first and second laws ideal gas processes vapor refrigeration cycles mixtures and solutions For students in engineering physics and chemistry **Power and the Engineer** ,1949 Engineering Fundamentals for Professional Engineers' Examinations Lloyd M. Polentz,1987

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