

Guided Practice - Acids & Bases**Properties of Acids and Bases**

1. Match the following terms with their descriptions:

- | | | |
|------------------------|-------|--|
| a. Arrhenius Acid | _____ | the substance formed when a proton is added to |
| b. Arrhenius Base | _____ | the remaining substance when a proton is lost from an acid |
| c. Conjugate Acid | _____ | a substance that produces OH ⁻ in solution |
| d. Conjugate Base | _____ | a substance producing H ⁺ in solution |
| e. Bronsted-Lowry Acid | _____ | a proton acceptor |
| f. Bronsted-Lowry Base | _____ | a proton donor |

2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are NOT conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.

- HSO_4^- ; SO_4^{2-}
- HBr ; BrO^-
- H_3PO_4 ; PO_4^{3-}
- H_2O ; HO^-

3. Complete the following table regarding acid strength.

| Property | Strong Acid | Weak Acid |
|---|-------------|-----------|
| Definition | | |
| Describe Dissociation | | |
| Species present in solution | | |
| Relative strength of conjugate base compared to water | | |
| Equation showing acid's reaction to water | | |

Acids Bases And Solutions Answer Key Lab35

NARAYAN CHANGDER



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