

5.6 Writing Formulas for Ionic Compounds Assignment

Instructions:

For this assignment, you are being asked to write formulas for Ionic Compounds and for Ionic Compounds containing Polyatomic Ions. Review the examples for each section and then write the formula for the identified compounds.

Writing Formulas for Ionic Compounds Assignment

Examples:

Write the formula of potassium bromide.

- K^{+1} Br^{-1}
- The sum of the charges is $+1 + (-1) = 0$, therefore, the formula is KBr

Write the formula of calcium bromide.

- Ca^{+2} Br^{-1}
- The sum of the oxidation numbers is $+2 + (-1) = +1$, therefore, we need a second bromide (-1) to obtain a proper formula.
- The new sum is $+2 + 2(-1) = 0$, therefore, the formula is $CaBr_2$.

A simpler way to understand how to write formulas is to use the crisscross & reduce method. If there is only one atom of an element in the compound, you do not need to put the number 1 behind it. The following are examples:



Questions 1-6.

Write formulas for the following compounds:

- | | | | |
|-------------------------|---------|-------------------------|-----------|
| 1. sodium and fluorine | NaF | 4. aluminum and sulfur | Al_2S_3 |
| 2. magnesium and iodine | MgI_2 | 5. hydrogen and oxygen | H_2O |
| 3. barium and oxygen | BaO | 6. lithium and nitrogen | Li_3N |

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Table of Contents Assignment 5 Ionic Compounds

1. Understanding the eBook Assignment 5 Ionic Compounds
 - The Rise of Digital Reading Assignment 5 Ionic Compounds
 - Advantages of eBooks Over Traditional Books
2. Identifying Assignment 5 Ionic Compounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Assignment 5 Ionic Compounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Assignment 5 Ionic Compounds

- Personalized Recommendations
 - Assignment 5 Ionic Compounds User Reviews and Ratings
 - Assignment 5 Ionic Compounds and Bestseller Lists
5. Accessing Assignment 5 Ionic Compounds Free and Paid eBooks
 - Assignment 5 Ionic Compounds Public Domain eBooks
 - Assignment 5 Ionic Compounds eBook Subscription Services
 - Assignment 5 Ionic Compounds Budget-Friendly Options
 6. Navigating Assignment 5 Ionic Compounds eBook Formats
 - ePub, PDF, MOBI, and More
 - Assignment 5 Ionic Compounds Compatibility with Devices
 - Assignment 5 Ionic Compounds Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Assignment 5 Ionic Compounds
 - Highlighting and Note-Taking Assignment 5 Ionic Compounds
 - Interactive Elements Assignment 5 Ionic Compounds
 8. Staying Engaged with Assignment 5 Ionic Compounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Assignment 5 Ionic Compounds
 9. Balancing eBooks and Physical Books Assignment 5 Ionic Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Assignment 5 Ionic Compounds
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Assignment 5 Ionic Compounds
 - Setting Reading Goals Assignment 5 Ionic Compounds
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Assignment 5 Ionic Compounds

- Fact-Checking eBook Content of Assignment 5 Ionic Compounds
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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