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displacement 2 $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$ double displacement 3 $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ combustion 4 $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$
single displacement 5 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ decomposition 6 $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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including the proper phase labels chemical equations can also be used to represent physical processes write a chemical

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 $4\text{Na} + \text{S} + 2\text{Cl}_2 \rightarrow 2\text{Na}_2\text{S} + 2\text{Cl}_2$

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20 2 src12 2 agn03 sr no $2\text{AgCl} + 3\text{LiBr} + 2\text{LiF} + 2\text{NaF} + 5\text{F}_2 + 2\text{Br}_2$ b balance the chemical reactions and identify each type

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web 1 $3\text{NaBr} + \text{H}_3\text{PO}_4 + \text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement 2 $3\text{Ca} + \text{OH}_2 + 1\text{Al}_2\text{SO}_4 + 3\text{CaSO}_4 + 2\text{Al} + \text{OH}_3$ type of reaction

double displacement 3 $3\text{Mg} + 1\text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$ type of reaction single displacement 4 $1\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ type of reaction

combustion 5 $2\text{PbSO}_4 + 2\text{PbSO}_3 + 1\text{O}_2$ type of reaction

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differentiate between a physical and chemical change physical does not change the identity of the matter can be changed

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