

A Water Quality Field Study

(adapted from Middle School Science with Verneir)

Introduction

When finding water quality, there are many measurements that can be made. In this experiment, you will measure water temperature, conductivity, pH, and dissolved oxygen.

Temperature

Water temperatures in streams can range from 0°C in the winter to above 30°C in the summer. Cooler water in a stream is generally considered healthier than warmer water. Problems generally occur when changes in water temperature are noted along one stream on the same day.

Conductivity

Conductivity values in lakes and streams are typically found to be in the range of 100 to 500 $\mu\text{S}/\text{cm}$. In areas of especially hard water or high salinity, conductivity values may be as high as 1000 $\mu\text{S}/\text{cm}$. Drinking water usually has conductivity in the 50 to 1000 $\mu\text{S}/\text{cm}$ range. Some sample data are listed in Table 2.

Conductivity is a measurement of the ability of an aqueous solution to carry an electrical current. An ion is an atom of an element that has gained or lost an electron which will create a negative or positive state. For example, sodium chloride (table salt) consists of sodium ions (Na^+) and chloride ions (Cl^-) held together in a crystal. In water it breaks apart into an aqueous solution of sodium and chloride ions. This solution will conduct an electrical current.

There are several factors that determine the degree to which water will carry an electrical current.

- 1) the concentration or number of ions;
- 2) mobility of the ion;
- 3) oxidation state (valence) and;
- 4) temperature of the water.

Used as a measure of the mineral or other ionic concentration. Conductivity is a measure of the purity of water or the concentration of ionized chemicals in water.

pH

The best pH range for most aquatic organisms is pH 6.5 to 8.2. The pH values of streams and lakes are usually between pH 7 and 8. Hard water will often have pH values between 7.5 and 8.5.

It is used to measure the acidity of liquids represented by a scale of numbers from 0 to 14. The pH of pure water is 7.0, which is neither acidic nor basic. When the pH is less than 7.0 it is identified as acidic; and at a pH greater than 7.0, it is considered basic or alkaline.

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