

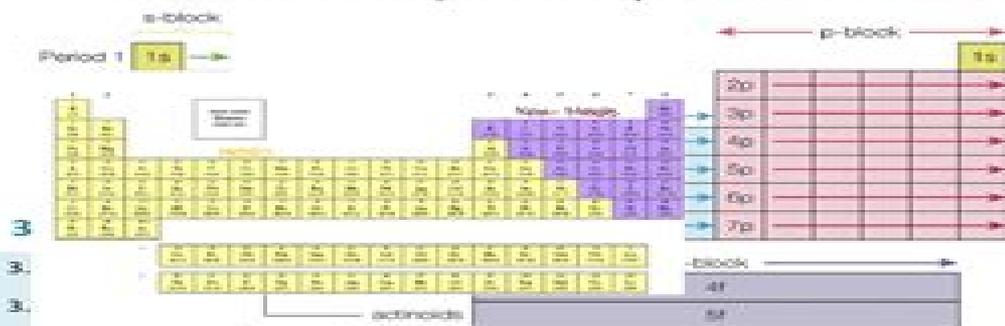
Topic 3: Periodicity

3.1 Periodic Table

- 3.1.1 The periodic table is arranged into four blocks associated with the four sub-levels – s, p, d, and f
- 3.1.2 The periodic table consists of groups (vertical columns) and periods (horizontal rows)
- 3.1.3 The period number (n) is the outer energy level that is occupied by electrons
- 3.1.4 The number of the principal energy level and the number of the valence electrons in an atom can be deduced from its position on the periodic table
- 3.1.5 The periodic table shows the position of metals, non-metals and metalloids
- 3.1.6 Deduction of the electron configuration of an atom from the element's position on the periodic table, and vice versa

Periodic Table Arrangement

- The periodic table is arranged into 4 blocks associated with their last orbital: s, p, d and f
- Elements in the periodic table are arranged by increasing atomic number
- **Group:** A vertical column of elements
 - Group number is the same as the number of valence electrons
- **Period:** A horizontal row of elements
 - Period number is the same number of shells in the atom
 - All shells except the outer shell will be full
- There is a division between metals and non-metals. Metals are on the left and non-metals are on the right
- The farther sub-level occupied is found by the location of the element (which block it is in):



- The location of the following families of the periodic table must be known: alkali metals, alkali earth metals, halogens, noble gases, transition metals, lanthanides and actinoids

- 3.2.1 Atomic radius, ionic radius
- 3.2.2 Electronegativity, ionization energy
- 3.2.3 Oxides change from basic through amphoteric to acidic across a period
- 3.2.4 Prediction and explanation of the metallic and non-metallic behavior of an element based on its position in the periodic table
- 3.2.5

Atomic Radius – The distance from the nucleus to the outermost electron

Ionic Radius – The distance from the nucleus to the outermost electrons in an ion

First ionization energy – The energy required to remove one mole of electrons from one mole of gaseous atoms. This is shown by:

Electron affinity – The energy released when one mole of an electron is added to one mole gaseous atoms. This is shown by:

Electronegativity – A measure of the attraction an atom has for a shared pair of electrons in a covalent bond

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Table of Contents Chemistry Topic 3 Periodic Table Outline Answers

1. Understanding the eBook Chemistry Topic 3 Periodic Table Outline Answers
 - The Rise of Digital Reading Chemistry Topic 3 Periodic Table Outline Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Topic 3 Periodic Table Outline Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Topic 3 Periodic Table Outline Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Topic 3 Periodic Table Outline Answers
 - Personalized Recommendations
 - Chemistry Topic 3 Periodic Table Outline Answers User Reviews and Ratings

- Chemistry Topic 3 Periodic Table Outline Answers and Bestseller Lists
- 5. Accessing Chemistry Topic 3 Periodic Table Outline Answers Free and Paid eBooks
 - Chemistry Topic 3 Periodic Table Outline Answers Public Domain eBooks
 - Chemistry Topic 3 Periodic Table Outline Answers eBook Subscription Services
 - Chemistry Topic 3 Periodic Table Outline Answers Budget-Friendly Options
- 6. Navigating Chemistry Topic 3 Periodic Table Outline Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Topic 3 Periodic Table Outline Answers Compatibility with Devices
 - Chemistry Topic 3 Periodic Table Outline Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Topic 3 Periodic Table Outline Answers
 - Highlighting and Note-Taking Chemistry Topic 3 Periodic Table Outline Answers
 - Interactive Elements Chemistry Topic 3 Periodic Table Outline Answers
- 8. Staying Engaged with Chemistry Topic 3 Periodic Table Outline Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Topic 3 Periodic Table Outline Answers
- 9. Balancing eBooks and Physical Books Chemistry Topic 3 Periodic Table Outline Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Topic 3 Periodic Table Outline Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Topic 3 Periodic Table Outline Answers
 - Setting Reading Goals Chemistry Topic 3 Periodic Table Outline Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Topic 3 Periodic Table Outline Answers
 - Fact-Checking eBook Content of Chemistry Topic 3 Periodic Table Outline Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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