

Section One: Multiple-choice

(50 marks)

1. D	2. C	3. B	4. C	5. D
6. C	7. C	8. D	9. B	10. C
11. D	12. B	13. B	14. C	15. A
16. All	17. B	18. D	19. A	20. C
21. C	22. B	23. C	24. D	25. C

Section Two: Short answer

(70 marks)

Question 26

(4 marks)

- (a) Some hydrochloric acid solution is mixed with solid sodium carbonate. (2 marks)

A white solid dissolves in a colourless solution, producing a colourless and odourless gas. (2)

(*Must have two observations for both marks).

- (b) Some solid copper (II) hydroxide is mixed with a dilute nitric acid solution. (2 marks)

A blue solid dissolves in a colourless solution to produce a blue solution. (2)

(*Must have two observations for both marks).

Question 27

(6 Marks)

- (a) Write balanced chemical equations showing carbon dioxide reacting with water to form carbonic acid, and then the two successive ionisation reactions that carbonic acid undergo in water. (3 marks)



- (b) Identify a conjugate acid-base pair in this reaction, and explain why it is classified as a Bronsted – Lowry acid-base reaction. (3 marks)

Conjugate A/B pair = $\text{CO}_3^{2-} / \text{HCO}_3^-$ (1) *Also accept $\text{HCO}_3^- / \text{CO}_3^{2-}$

This equation is classified as a Bronsted – Lowry acid-base reaction because in the forward reaction, H_2O donates a proton, thus acting as a B-L acid, (1)
while CO_3^{2-} accepts a proton, thus acting as a B-L base. (1)

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