

Chapter 1 - Introduction

Classifications of Matter – 1.3

- * **Matter:** anything that occupies space and has mass
- * **Chemistry:** the study of matter and its change
- * **Atom:** smallest distinctive unit
- * **Molecule:** two or more atoms together
- * **Substance:** group of identical atoms or molecules
 - o **Element:** atoms are identical
 - o **Compound:** contains different atoms
- * **Homogenous Mixture:** uniform, seems as one thing
- * **Heterogeneous Mixture:** non-uniform, able to see the elements



Physical and Chemical Properties of Matter – 1.4

- * **Physical Property:** measurable, describes system state
 - o Melting point, solubility, density, color
 - o "Air is a gas," "boiling point of water is 100 degrees," "Sulfur crystals are yellow"
- * **Chemical Property:** chemical reaction, compound cannot be restored without a change
 - o Burning, explosions, decomposition, rusting
 - o i.e. "Burning Hydrogen in Oxygen to form water," "milk turning sour"



Measurement – 1.5

- * **SI base units of Measurement:**
 - o Length – meter (m)
 - o Mass – kilogram (kg)
 - o Time – second (s)
 - o Temperature – Kelvin (K)
 - o Amount of a substance – mole (mol)
- * **Temperature Scales:**
 - o Kelvin: absolute temperature scale, $273^{\circ} \text{C} = 1 \text{K}$

Handling Number – 1.6

- * **Accuracy:** Correctness – How close the results were to where you intended them to be
- * **Precision:** Exactness – How close the results were in relation to other results / trials
- * **Scientific Notation:** Expresses a number as a product of a number between 1 and 10 and the appropriate power of 10
 - o All non-zeros are **significant** (1-9)
 - o A zero between non-zeros is **significant** (209)
 - o Final zeros to the right of the decimal are **significant** (10.0)
 - o Initial zeros are **not significant** (0.47)
 - o Final zeros in a number with no decimal **may or may not be significant** (9,500)
- * **Adding and subtracting:** same number of digits past the decimal point as the number used with the least precision – the least amount of significant figures
 - o If greater than 5, round up
- * **Multiplication and Division:** least number of significant figures
- * **Exact Numbers:** have no effect on significant figures – i.e. 60 seconds in a minute
- * **Volume (V):** Length cubed, $\text{mL} = 1 \text{cm}^3$

Chemistry 101 Study Guide

Lingjun Ying



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Table of Contents Chemistry 101 Study Guide

1. Understanding the eBook Chemistry 101 Study Guide
 - The Rise of Digital Reading Chemistry 101 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 101 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 101 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 101 Study Guide
 - Personalized Recommendations
 - Chemistry 101 Study Guide User Reviews and Ratings
 - Chemistry 101 Study Guide and Bestseller Lists

5. Accessing Chemistry 101 Study Guide Free and Paid eBooks
 - Chemistry 101 Study Guide Public Domain eBooks
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 - Chemistry 101 Study Guide Budget-Friendly Options
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 - Chemistry 101 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 101 Study Guide
 - Highlighting and Note-Taking Chemistry 101 Study Guide
 - Interactive Elements Chemistry 101 Study Guide
8. Staying Engaged with Chemistry 101 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 101 Study Guide
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 101 Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 101 Study Guide
 - Setting Reading Goals Chemistry 101 Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 101 Study Guide
 - Fact-Checking eBook Content of Chemistry 101 Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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