

**CHAPTER 9 REVIEW***Stoichiometry***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ The coefficients in a chemical equation represent the
  - (a) masses in grams of all reactants and products.
  - (b) relative number of moles of reactants and products.
  - (c) number of atoms of each element in each compound in a reaction.
  - (d) number of valence electrons involved in a reaction.
  
2. \_\_\_\_\_ Which of the following would not be studied within the topic of stoichiometry?
  - (a) the mole ratio of Al to Cl in the compound aluminum chloride
  - (b) the mass of carbon produced when a known mass of sucrose decomposes
  - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
  - (d) the amount of energy required to break the ionic bonds in  $\text{CaF}_2$
  
3. \_\_\_\_\_ A balanced chemical equation allows you to determine the
  - (a) mole ratio of any two substances in the reaction.
  - (b) energy released in the reaction.
  - (c) electron configuration of all elements in the reaction.
  - (d) reaction mechanism involved in the reaction.
  
4. \_\_\_\_\_ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
  - (a) reaction sequence.
  - (b) bond energy.
  - (c) mole ratio.
  - (d) element proportion.
  
5. Given the reaction represented by the following unbalanced equation:  $\text{N}_2\text{O}(g) + \text{O}_2(g) \rightarrow \text{NO}_2(g)$ 
  - a. Balance the equation.  
\_\_\_\_\_  
\_\_\_\_\_
  - b. What is the mole ratio of  $\text{NO}_2$  to  $\text{O}_2$ ?  
\_\_\_\_\_
  - c. If 20.0 mol of  $\text{NO}_2$  form, how many moles of  $\text{O}_2$  must have been consumed?  
\_\_\_\_\_
  - d. Twice as many moles of  $\text{NO}_2$  form as moles of  $\text{N}_2\text{O}$  are consumed. True or False?  
\_\_\_\_\_
  - e. Twice as many grams of  $\text{NO}_2$  form as grams of  $\text{N}_2\text{O}$  are consumed. True or False?  
\_\_\_\_\_

# Chapter 9 Mixed Review Stoichiometry Answers

**Sebastian Brünink**



## Chapter 9 Mixed Review Stoichiometry Answers:

**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001     *Holt Chemistry* Ralph Thomas Myers, 2004     **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall, Donald J. DeCoste, 2000     **Basic Chemistry** Steven S Zumdahl, 1996     *Thermodynamic Data* Surendra K. Saxena, 2012-12-06 With the rapid development of fast processors the power of a mini super computer now exists in a lap top box Quite sophisticated techniques are becoming accessible to geoscientists thus making disciplinary boundaries fade Chemists and physicists are no longer shying away from computational mineralogical and material science problems too complicated to handle Geoscientists are willing to delve into quantitative physico chemical methods and open those black boxes they had shunned for several decades but with which had learned to live I am proud to present yet another volume in this series which is designed to break the disciplinary boundaries and bring the geoscientists closer to their chemist and physicist colleagues in achieving a common goal This volume is the result of an international collaboration among many physical geochemists chemists physicists and geologists aiming to understand the nature of material The book has one common theme namely how to determine quantitatively through theory the physico chemical parameters of the state of a solid or fluid     *The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB     [Chemical Engineering and Mining Review](#), 1915     *Applied Mechanics Reviews*, 1986     *Chemical Principles* Peter Atkins, Loretta Jones, 2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding     **Introductory Chemistry** Steven S. Zumdahl, 1996     *Chemical Analysis* Kenneth A.

Rubinson,1987 *General, Organic, and Biochemistry* William Henry Brown,Elizabeth P. Rogers,1980 Industrial & Engineering Chemistry ,1926 *CHEMISTRY* SILBERBERG,2003 **Chemistry** John S. Phillips,Cheryl Wistrom,2000 **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Battelle Technical Review** ,1953 **Theoretical Chemical Engineering Abstracts** ,1984 *Complete Preparation for the MCAT* Williams & Wilkins Review,1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT **A Complete Preparation for the MCAT** Aftab S. Hassan,James L. Flowers,1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress

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