

Stoichiometry Review

****YOU MUST SHOW ALL WORK AND YOUR ANSWERS MUST INCLUDE THE PROPER NUMBER OF SIG FIGS AND COMPLETE UNITS IN ORDER TO RECEIVE CREDIT FOR THE PROBLEM.****

BALANCE THE FOLLOWING EQUATIONS TO USE IN QUESTIONS 5 – 14:



PERFORM THE FOLLOWING STOICHIOMETRIC CALCULATIONS:

5. Zinc reacts with hydrochloric acid to produce zinc chloride and hydrogen. How many moles of HCl are required to produce 7.50 moles of ZnCl₂?

$$\frac{7.50 \text{ mol ZnCl}_2}{1 \text{ mol ZnCl}_2} \times \frac{2 \text{ mol HCl}}{1 \text{ mol ZnCl}_2} = 15.0 \text{ mol HCl}$$

6. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many grams of copper are required to form 250 g of silver?

$$\frac{250 \text{ g Ag}}{107.868 \text{ g Ag}} \times \frac{1 \text{ mol Ag}}{2 \text{ mol Ag}} \times \frac{1 \text{ mol Cu}}{1 \text{ mol Cu}} \times 63.546 \text{ g Cu} = 74 \text{ g Cu}$$

7. When aluminum is burned in excess oxygen, aluminum oxide is produced. How many grams of oxygen are required to produce 0.75 moles of Al₂O₃?

$$\frac{0.75 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}_2\text{O}_3} \times \frac{3 \text{ mol O}_2}{1 \text{ mol O}_2} \times 31.998 \text{ g O}_2 = 36 \text{ g O}_2$$

8. How many grams of iron(III) chloride are produced when 15.3 g of iron react with excess chlorine gas?

$$\frac{15.3 \text{ g Fe}}{55.845 \text{ g Fe}} \times \frac{1 \text{ mol Fe}}{2 \text{ mol Fe}} \times \frac{2 \text{ mol FeCl}_3}{1 \text{ mol FeCl}_3} \times 162.204 \text{ g FeCl}_3 = 44.4 \text{ g FeCl}_3$$

9. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many moles of silver will be produced from 3.65 moles of silver nitrate?

$$\frac{3.65 \text{ mol AgNO}_3}{2 \text{ mol AgNO}_3} \times \frac{2 \text{ mol Ag}}{1 \text{ mol AgNO}_3} = 3.65 \text{ mol Ag}$$

Chapter 8 Stoichiometry Review Sheet Answers

R Sandford



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