

## Chapter 2 Minerals

### Learning Objectives

After reading this chapter, completing the exercises within it, and answering the questions at the end, you should be able to:

- Describe the nature of atoms and their constituents, particularly the behaviour of electrons and the formation of ions.
- Apply your understanding of atoms to explain bonding within minerals.
- Describe mineral lattices and explain how they influence mineral properties.
- Categorise minerals into groups based on their compositions.
- Describe a silica tetrahedron and the ways in which tetrahedra combine to make silicate minerals.
- Differentiate between ferromagnesian and other silicate minerals.
- Explain some of the mechanisms of mineral formation.
- Describe some of the important techniques for identifying minerals.

Minerals are all around us: the graphite in your pencil, the salt on your table, the plaster on your walls, and the trace amounts of gold in your computer. Minerals can be found in a wide variety of consumer products including paper, medicine, processed foods, cosmetics, electronic devices, and many more. And of course, everything made of metal is also derived from minerals.

As defined in Chapter 1, a mineral is a naturally occurring combination of specific elements arranged in a particular repeating three-dimensional structure (Figure 1.4.1).

“Naturally occurring” implies that minerals are not artificially made. Many minerals (e.g., diamond) can be made in laboratories, but if they can also occur naturally, they still qualify as minerals.

“Specific elements” means that most minerals have a specific chemical formula or composition. The mineral pyrite, for example, is  $\text{FeS}_2$  (two atoms of sulfur for each atom of iron), and any significant departure from that would make it a different mineral. But many minerals can have variable compositions within a specific range. The mineral olivine, for example, can range all the way from  $\text{Fe}_2\text{SiO}_4$  to  $\text{FeMg}_2\text{SiO}_4$  to  $\text{Mg}_2\text{SiO}_4$ . Intervening compositions are written as  $(\text{Fe},\text{Mg})_2\text{SiO}_4$  meaning that Fe and Mg can be present in any proportion, and that there are two of them for each Si present. This type of substitution is known as **solid solution**.

Most important of all, a mineral has a specific “repeating three-dimensional structure” or “lattice,” which is the way in which the atoms are arranged. We’ve already seen in Chapter 1 how sodium and chloride atoms in halite alternate in a regular pattern. That happens to be about the simplest mineral lattice of all; most mineral lattices are much more complicated, as we’ll see.

Some substances that we think must be minerals are not because they lack that repeating

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## **Table of Contents Chapter 2 Minerals**

1. Understanding the eBook Chapter 2 Minerals
  - The Rise of Digital Reading Chapter 2 Minerals
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 2 Minerals
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 2 Minerals
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 2 Minerals
  - Personalized Recommendations
  - Chapter 2 Minerals User Reviews and Ratings
  - Chapter 2 Minerals and Bestseller Lists
5. Accessing Chapter 2 Minerals Free and Paid eBooks
  - Chapter 2 Minerals Public Domain eBooks

- Chapter 2 Minerals eBook Subscription Services
- Chapter 2 Minerals Budget-Friendly Options
- 6. Navigating Chapter 2 Minerals eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 2 Minerals Compatibility with Devices
  - Chapter 2 Minerals Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 2 Minerals
  - Highlighting and Note-Taking Chapter 2 Minerals
  - Interactive Elements Chapter 2 Minerals
- 8. Staying Engaged with Chapter 2 Minerals
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 2 Minerals
- 9. Balancing eBooks and Physical Books Chapter 2 Minerals
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 2 Minerals
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 2 Minerals
  - Setting Reading Goals Chapter 2 Minerals
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 2 Minerals
  - Fact-Checking eBook Content of Chapter 2 Minerals
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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