

Notes Ch. 4 and 25: Atomic Structure and Nuclear Chemistry
History and Structure of the Nuclear Atom

The Atom

- smallest particle of an element that retains all properties of the element

I. Early Models of the Atom

A. Democritus (460 B.C. – 370 B.C.)

- first to suggest the existence of _____
- **believed** atoms were indivisible and indestructible

B. Dalton

- Atoms of the same element are _____
- Each element is unique
- Atoms combine in fixed ratios to form compounds

C. Thomson

- Discovered _____
- Suggested atom looked like plum pudding (or a blueberry muffin or chocolate chip cookie) with electrons evenly distributed throughout positive sphere

D. Rutherford

- Conducted gold foil experiment
- Proposed atom is mostly empty space
- Concluded all positive charge and mass is concentrated in small region called _____

E. Bohr

- depicts the atom as a small, positively charged nucleus
- surrounded by electrons that travel in circular _____ around the nucleus

F. Schrödinger

- model allowed the electron to occupy three-dimensional space like an electron " _____ "
- Used a mathematical equation to show a model of electrons as waves

II. More about the Atom

A. Size

- Teeny tiny
- observable with instruments such as a scanning tunneling _____

B. Parts (called Subatomic Particles)

1. Protons: p^+

- charge = positive +1
- mass = 1

2. Electrons: e^-

- charge = negative -1
- mass ~ 0

3. Neutrons: n^0

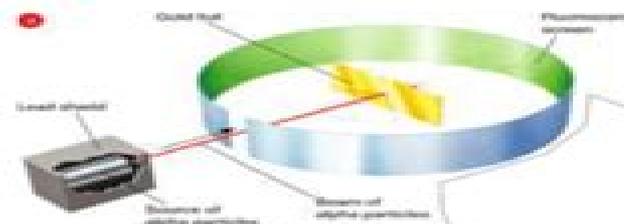
- charge = zero
- mass = 1

C. Location of Parts

- **Nucleus:** center of atom
 - contains protons and neutrons
 - has a positive charge
 - contains almost ALL of the _____ of the atom
- **Outside nucleus:**
 - contains _____
 - has a negative charge
 - occupies almost ALL of the volume of the atom

D. Atomic Charge

- _____ are NEUTRAL
- # of protons = # of electrons
(ALWAYS if you're talking about an ATOM)



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A Loxley



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