

Question number	Answer	Marks	Guidance
1 (a)	$n(\text{NaOH}) = 0.125 \times 22.40/1000 = 2.80 \times 10^{-3} \text{ mol}$	B1	
1 (b)	$n(\text{NaOH}) = 0.5 \times 2.80 \times 10^{-3} = 1.40 \times 10^{-3} \text{ mol}$	B1	
1 (c)	$c = 1.40 \times 10^{-3} \times 1000/25.0 = 0.0560 \text{ mol dm}^{-3}$	B1	
2	N in NH_3 has been oxidised from -3 to 0 in N_2 Cu in CuO has been reduced from $+2$ to 0 in Cu	B1 B1	
3 (a) (i)	Mg has lost 2 electrons and has been oxidized Fe has lost 3 electrons and has been reduced	B1 B1	
3 (a) (ii)	iron(III) nitrate(V)	B1	
3 (b) (i)	Mn has been reduced from $+4$ in MnO_2 to $+2$ in MnCl_2	B2	
3 (b) (ii)	In HCl , Cl has oxidation number of -1 . Cl has been oxidised from -1 in HCl to 0 in Cl_2 Cl in HCl has an oxidation number of -1 which is unchanged in MnCl_2	B1 B1	
4 (a) (i)	For each equation, 1 mark for balanced equation, 1 mark for state symbols. $\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ $\text{CaO}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$ $\text{Ca}(\text{OH})_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$	B1 x 2 B1 x 2 B1 x 2	
4 (a) (ii)	An acid has been neutralised by a base to form water	B1	
4 (a) (iii)	$\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$	B1	
4 (b) (i)	Reduction is increase in oxidation number and gain of electrons Oxidation is decrease in oxidation number and loss of electrons	B1 B1	
4 (b) (ii)	$\text{Ca}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{g}) + \text{H}_2(\text{g})$	B2	
4 (b) (iii)	Ca has been oxidised from 0 to $+2$ H has been reduced from $1+$ to 0 .	B1 B1	
5 (a)	$\text{WO}_3 + 3\text{H}_2 \rightarrow \text{W} + 3\text{H}_2\text{O}$	B1	

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