

Chemistry: Molality

Directions: Using these formulas, solve the following problems.

$$\text{molality} = \text{mol solute} / \text{kg solvent}$$

$$1 \text{ amu} = 1 \text{ gram} / \text{mol}$$

$$\Delta T_{\text{freezing}} = (- \text{number of ions}) (\text{molality}) (k_f)$$

$$\Delta T_{\text{boiling}} = (\text{number of ions}) (\text{molality}) (k_b)$$

k_b and k_f are constants for each individual solvent.

Key



Problem 1: What is the molality of a solution containing 1.23 g of $\text{Hg}(\text{CN})_2$ dissolved in 1.37 g of methanol? Answer: 3.55 m

Problem 2: What is the molality of a solution containing 16.8 g of NaCl dissolved in 356 g of water? Answer: 0.806 m

Problem 3: How many grams of NaOH would be required to produce a 3.57 m solution with 473 g of water? Answer: 67.6 g

Problem 4: How many grams of KCl would you need to dissolve in 114 g of vinegar in order to prepare a 1.96 m solution? Answer: 16.6 g

Problem 5: Calculate the molality of 5.97 g NaCl dissolved in 471 g of water.
Answer: 0.217 m

Problem 6: Calculate the molality of 8.91 g Na_2CO_3 dissolved in 595 g of water.
Answer: 0.341 m

Problem 7: Calculate the molality of 8.44 g KBr dissolved in 112 g of water.
Answer: 0.639 m

Problem 8: What are the new boiling and freezing points for a solution with 581 g KBr dissolved in 845 grams of water? (k_b water = 0.512°C/m and k_f water = 1.86°C/m)
Answer: new freezing point = -13.4°C new boiling point = 106°C

Problem 9: What are the new boiling and freezing points for a solution with 391 g NaOH dissolved in 1029 grams of water? (k_b water = 0.512°C/m and k_f water = 1.86°C/m)
Answer: new freezing point = -36.3°C new boiling point = 107°C

Molality Answers Key

James P. Birk, Donald M. Baird



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Table of Contents Molality Answers Key

1. Understanding the eBook Molality Answers Key
 - The Rise of Digital Reading Molality Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molality Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molality Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molality Answers Key
 - Personalized Recommendations

- Molality Answers Key User Reviews and Ratings
- Molality Answers Key and Bestseller Lists
- 5. Accessing Molality Answers Key Free and Paid eBooks
 - Molality Answers Key Public Domain eBooks
 - Molality Answers Key eBook Subscription Services
 - Molality Answers Key Budget-Friendly Options
- 6. Navigating Molality Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molality Answers Key Compatibility with Devices
 - Molality Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molality Answers Key
 - Highlighting and Note-Taking Molality Answers Key
 - Interactive Elements Molality Answers Key
- 8. Staying Engaged with Molality Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molality Answers Key
- 9. Balancing eBooks and Physical Books Molality Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molality Answers Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molality Answers Key
 - Setting Reading Goals Molality Answers Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molality Answers Key
 - Fact-Checking eBook Content of Molality Answers Key

-
- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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