

# CHEM 103 Lab Exam 3 (100 out of 100) Questions and Verified Answers (Portage Learning)

## Scientific Calculator

### Question 1

5 / 10 pts

Answer the following two questions:

1. Why was the hydrate heated with the Fisher burner in part 2 of lab 3?
2. For what purpose was SEM carried out in lab 3?

1. The hydrate was heated with the Fisher burner in lab 3 to cause loss of water.
2. The SEM was carried out in lab 3 to detect the elemental composition.

2. The SEM was carried out in lab 3 to detect the elemental composition. (We did not use it for imaging)

### Question 2

10 / 10 pts

## Scientific Calculator

Answer the following question.

A 1.1000 gram hydrate sample chosen from  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ ,  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$  and  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  was heated and found to lose 0.6920 gram of  $\text{H}_2\text{O}$ . (1) Show the calculation of the %  $\text{H}_2\text{O}$  in the unknown hydrate sample. (2) Show the calculation of the %  $\text{H}_2\text{O}$  in each of the hydrate compounds and identify the unknown hydrate from the list.

Atomic weights: H = 1.008, O = 16.00. MWs:  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$  = 286.15,  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$  = 241.43,  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$  = 203.301 and  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  = 244.462

%  $\text{H}_2\text{O}$  in unknown =  $(0.6920 / 1.1000) \times 100 = 62.91\%$

%  $\text{H}_2\text{O}$  in  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$  =  $180.16 / 286.15 \times 100 = 62.96\%$

unknown is  **$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$**

### Question 3

10 / 10 pts

# Chem 110 Lab Questions And Answers

**Lei Huang**



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### **Table of Contents Chem 110 Lab Questions And Answers**

1. Understanding the eBook Chem 110 Lab Questions And Answers
  - The Rise of Digital Reading Chem 110 Lab Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chem 110 Lab Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chem 110 Lab Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 110 Lab Questions And Answers
  - Personalized Recommendations
  - Chem 110 Lab Questions And Answers User Reviews and Ratings
  - Chem 110 Lab Questions And Answers and Bestseller Lists

5. Accessing Chem 110 Lab Questions And Answers Free and Paid eBooks
  - Chem 110 Lab Questions And Answers Public Domain eBooks
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  - Chem 110 Lab Questions And Answers Budget-Friendly Options
6. Navigating Chem 110 Lab Questions And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chem 110 Lab Questions And Answers Compatibility with Devices
  - Chem 110 Lab Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chem 110 Lab Questions And Answers
  - Highlighting and Note-Taking Chem 110 Lab Questions And Answers
  - Interactive Elements Chem 110 Lab Questions And Answers
8. Staying Engaged with Chem 110 Lab Questions And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chem 110 Lab Questions And Answers
9. Balancing eBooks and Physical Books Chem 110 Lab Questions And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chem 110 Lab Questions And Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chem 110 Lab Questions And Answers
  - Setting Reading Goals Chem 110 Lab Questions And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem 110 Lab Questions And Answers
  - Fact-Checking eBook Content of Chem 110 Lab Questions And Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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