

Name \_\_\_\_\_

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/Atp	VSEPR			
			Electron domain geometry	Molecular geometry	Hybridization of central atom	Polar ?
CO <sub>2</sub>	<p>Valence 24</p> $\text{O}=\text{C}=\text{O}$ <p>2 charge clouds</p>	0	Linear 180°	Linear 	sp	NP
BF <sub>3</sub>	<p>Valence 32</p> $\text{F}=\text{B}-\text{F}$	0	Trigonal planar 120°		sp <sup>2</sup>	NP
SnCl <sub>2</sub>	<p>Valence 24</p> $\text{Cl}=\text{Sn}-\text{Cl}$	3	Bent 120°		sp <sup>2</sup>	P
BF <sub>4</sub> <sup>-</sup>	<p>Valence 32</p> $\left[ \text{F}-\text{B}-\text{F} \right]^-$	0	Tetrahedral 109.5°		sp <sup>3</sup>	P
AsH <sub>3</sub>	<p>Valence 24</p> $\text{H}-\text{As}-\text{H}$	4	Trigonal pyramidal 109.5°		sp <sup>3</sup>	P
H <sub>2</sub> S	<p>Valence 14</p> $\text{H}-\text{S}-\text{H}$	2	Bent 109.5°		sp <sup>3</sup>	NP
NH <sub>2</sub> <sup>-</sup>	<p>Valence 10</p> $\left[ \text{H}-\text{N}-\text{H} \right]^-$	2	Bent 109.5°		sp <sup>3</sup>	P
PCl <sub>4</sub>	<p>Valence 40</p> $\text{F}-\text{P}-\text{Cl}$	0	Trigonal bipyre 120 + 90		sp <sup>3</sup> d	P

# Molecular Shapes Lab Activity Answers

**Padhraic Smyth**



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## **Table of Contents Molecular Shapes Lab Activity Answers**

1. Understanding the eBook Molecular Shapes Lab Activity Answers
  - The Rise of Digital Reading Molecular Shapes Lab Activity Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Shapes Lab Activity Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molecular Shapes Lab Activity Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Shapes Lab Activity Answers
  - Personalized Recommendations
  - Molecular Shapes Lab Activity Answers User Reviews and Ratings
  - Molecular Shapes Lab Activity Answers and Bestseller Lists
5. Accessing Molecular Shapes Lab Activity Answers Free and Paid eBooks

- Molecular Shapes Lab Activity Answers Public Domain eBooks
  - Molecular Shapes Lab Activity Answers eBook Subscription Services
  - Molecular Shapes Lab Activity Answers Budget-Friendly Options
6. Navigating Molecular Shapes Lab Activity Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Molecular Shapes Lab Activity Answers Compatibility with Devices
  - Molecular Shapes Lab Activity Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Molecular Shapes Lab Activity Answers
  - Highlighting and Note-Taking Molecular Shapes Lab Activity Answers
  - Interactive Elements Molecular Shapes Lab Activity Answers
8. Staying Engaged with Molecular Shapes Lab Activity Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molecular Shapes Lab Activity Answers
9. Balancing eBooks and Physical Books Molecular Shapes Lab Activity Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molecular Shapes Lab Activity Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Molecular Shapes Lab Activity Answers
- Setting Reading Goals Molecular Shapes Lab Activity Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molecular Shapes Lab Activity Answers
- Fact-Checking eBook Content of Molecular Shapes Lab Activity Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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