



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
- Write the formula for the following compounds:
  - hydrogen chloride
  - hydrogen sulfide
  - hydrogen cyanide
  - hydrogen peroxide
  - hydrogen selenide
  - hydrogen telluride
- Write the formula for the following compounds:
  - hydrogen bromide
  - hydrogen iodide
  - hydrogen fluoride
  - hydrogen phosphide
  - hydrogen selenide
  - hydrogen telluride
- Write the formula for the following compounds:
  - hydrogen selenide
  - hydrogen telluride
  - hydrogen fluoride
  - hydrogen phosphide
  - hydrogen selenide
  - hydrogen telluride

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride.  $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromide.  $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodide.  $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide.  $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen sulfide.  $\text{H}-\text{S}-\text{H}$
- Draw the Lewis structure for hydrogen selenide.  $\text{H}-\text{Se}-\text{H}$
- Draw the Lewis structure for hydrogen telluride.  $\text{H}-\text{Te}-\text{H}$



### SECTION 8.3 BONDING IN THE CARBON

- Write the Lewis structure for ethane.  $\text{C}_2\text{H}_6$
- Write the Lewis structure for ethene.  $\text{C}_2\text{H}_4$
- Write the Lewis structure for ethyne.  $\text{C}_2\text{H}_2$

# Chapter 8 Covalent Bonding Practice Problems Answers

**Richard Bailey**



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## **Table of Contents Chapter 8 Covalent Bonding Practice Problems Answers**

1. Understanding the eBook Chapter 8 Covalent Bonding Practice Problems Answers
  - The Rise of Digital Reading Chapter 8 Covalent Bonding Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Covalent Bonding Practice Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Practice Problems Answers
  - Personalized Recommendations
  - Chapter 8 Covalent Bonding Practice Problems Answers User Reviews and Ratings
  - Chapter 8 Covalent Bonding Practice Problems Answers and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Practice Problems Answers Free and Paid eBooks
  - Chapter 8 Covalent Bonding Practice Problems Answers Public Domain eBooks
  - Chapter 8 Covalent Bonding Practice Problems Answers eBook Subscription Services
  - Chapter 8 Covalent Bonding Practice Problems Answers Budget-Friendly Options

6. Navigating Chapter 8 Covalent Bonding Practice Problems Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 8 Covalent Bonding Practice Problems Answers Compatibility with Devices
  - Chapter 8 Covalent Bonding Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Practice Problems Answers
  - Highlighting and Note-Taking Chapter 8 Covalent Bonding Practice Problems Answers
  - Interactive Elements Chapter 8 Covalent Bonding Practice Problems Answers
8. Staying Engaged with Chapter 8 Covalent Bonding Practice Problems Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 8 Covalent Bonding Practice Problems Answers
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Practice Problems Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Practice Problems Answers
  - Setting Reading Goals Chapter 8 Covalent Bonding Practice Problems Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Practice Problems Answers
  - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Practice Problems Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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