

## Chapter 6 Thermochemistry

### Homework Problems

- The enthalpy change for the fermentation (without oxygen,  $O_2$  gas) reaction of glucose to alcohol and carbon dioxide is  $-67$  kJ.
  - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
  - Is the reaction exothermic or endothermic?
  - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is  $0.71$  J/°C·g. Calculate the energy required to:
  - Raise the temperature of  $1.0$  mol graphite by  $1.0$ °C.
  - Raise the temperature of  $850$  g graphite by  $150$ °C.
  - Raise the temperature of  $75$  kg graphite from  $294$  K to  $348$  K.
- In a coffee-cup calorimeter,  $100.0$  mL of  $1.0$  M NaOH and  $100.0$  mL of  $1.0$  M HCl are mixed. Both solutions were originally at  $24.6$ °C. After the reaction, the final temperature is  $31.3$ °C. Assume that all the solutions have a density of  $1.0$  g/mL and a specific heat capacity of  $4.18$  J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ( $C_{15}H_{26}O$ ) has an energy of combustion of  $-5903.6$  kJ/mol. A sample of camphor with a mass of  $0.1204$  g is burned in a bomb calorimeter, the temperature increases by  $2.28$ °C. Calculate the heat capacity of the calorimeter.  
A  $0.1964$  g sample of quinone ( $C_6H_4O_2$ ) is burned in a bomb calorimeter that has a heat capacity of  $1.56$  kJ/°C. The temperature of the calorimeter increases by  $3.2$ °C. Calculate the energy of combustion of quinone per gram and per mole.

# Chapter 6 Thermochemistry Homework Problems

**TD Snyder**



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