

Solutions Manual to Accompany  
**Basic Principles and Calculations  
in Chemical Engineering**

Seventh Edition

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PRENTICE HALL  
Professional Technical Reference  
Upper Saddle River, New Jersey 07458

**TYPICAL EXAMS FOR A ONE SEMESTER COURSE**  
(Scheduled in the evening to avoid time constraints on students)

**Exam No. 1**  
(Open Book, 1 1/2 hours)

**PROBLEM 1 (25%)**

Hydrogen can be separated from natural gas by diffusion through a metal tube. The rate of separation is given by

$$N = 2 \pi DpR$$

where  
 $N$  = rate of transport of  $H_2$  from the tube, g moles/decade (of length of tube)  
 $D$  = diffusion coefficient  
 $p$  = molar density of  $H_2$ , g moles/cm<sup>3</sup>  
 $R$  = log mass ratio of tubes,  $r_2 - r_1 / r$  or  $\left(\frac{r_2}{r_1}\right)$ , with  $r$  in cm.

What are the units of  $D$ ?

**PROBLEM 2 (25%)**

A pellet of bones weighing 10 tons is dropped from a lift truck from a height of 10 feet. The maximum velocity that pellet attains before hitting the ground is 8 ft/sec. How much kinetic energy does the pellet have in (Joules) at this velocity?

**PROBLEM 3 (25%)**

The specific gravity of a fuel oil is 0.82. What is the density of the oil in lb/ft<sup>3</sup>? Show all units.

**PROBLEM 4 (25%)**

Sulfur trioxide ( $SO_3$ ) can be absorbed in sulfuric acid solution to form more concentrated sulfuric acid. If the gas to be absorbed contains 50%  $SO_3$ , 41%  $N_2$ , 7%  $SO_2$ , and 1%  $O_2$ , how many parts per million of  $O_2$  are there in the gas? (a) What is the composition of the gas on a  $H_2$  free basis?

**PROBLEM 5 (25%)**

You have 100 kilograms of gas of the following composition:

$CH_4$	30%
$H_2$	20%
$N_2$	50%

What is the average molecular weight of this gas?

**PROBLEM 6 (25%)**

If the heat capacity of a substance is 3.32 J/g(°C) and its molecular weight is 75.4, what is its heat capacity in

(a)	Btu(°F)
(b)	kJ(kg°C)
(c)	kJ(gmol°C)

**PROBLEM 7 (25%)**

A rock containing 100%  $BaSO_4$  is heated with coke (94% C, 6% ash) and the composition of the product is  $BaSO_4$  (11.5%),  $BaS$  (72.9%), C (13.9%), ash (2.2%). The reaction is



Calculate the percent excess reactant, and the degree of completion of the reaction.

**PROBLEM 8 (25%)**

A gas cylinder to which is attached an Bourdon gauge appears to be at a pressure of 21.38 in.  $H_2O$  at 70°F. The barometer reads 141.8 inHg. A student claims that the pressure in the tank is 1.3 psia, but another student points out that this is impossible - the pressure is really 28.2 psia. Can 1.3 psia be correct? Explain and show calculations to back up your explanation.

# Solution Basic Principles Himmelblau Ed 6

**Brendan G. Carr**



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