

# AP CHEMISTRY FORMULA SHEET 2025

Equations	Constants	Standard Reduction Potentials	
$PV = nRT$ $\ln[A] = -kt + \ln[A]_0$ $1/[A] = kt + 1/[A]_0$ $A = abc$ $\Delta G = -\Delta H - T\Delta S$ $q = mc\Delta T$ $q = -q$ $E = hy$ $pH = -\log[H^+]$ $pOH = -\log[OH^-]$ $pH + pOH = 14$ $K_a = [H^+][A^-]/[HA]$ $K_w = [H^+][OH^-] = 1.0 \times 10^{-14}$ $pK_a = -RT \ln K$ $pH = pK_a + \log [A^-]/[HA]$	$R = 0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K}$ $0.314 \text{ J/mol}\cdot\text{K}$ $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$ $F = 96,485 \text{ C/mol}\cdot\text{e}^-$ $c = 3.00 \times 10^8 \text{ m/s}$ $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$ $P = 1 \text{ atm}$ $T = 298 \text{ K (25}^\circ\text{C)}$	$\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$ $\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$ $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$ $2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$ $\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$	$\text{Cu}^{2+} \rightarrow \text{Cu} \quad E$ $\text{Zn}^{2+} \rightarrow \text{Zn} \quad 0.34$ $\text{Cl}_2 \rightarrow 2\text{Cl}^- \quad 1.36$ $\rightarrow \text{K} \quad 0.83$ $\text{Mg}^{2+} \rightarrow \text{Mg} \quad 2.37$
		Thermochemistry Data	
	<b>Thermochemistry Data</b> $\Delta H_f^\circ$ : kJ/mol Bond energies: kJ/mol Specific heat capacity, $C_p$ : J/g $^\circ\text{C}$	$\Delta H^\circ$ : kJ/mol Bond energies: kJ/mol Specific heat capacity, $C_p$ : J/g $^\circ\text{C}$	$\text{NO}_3^-$ $\text{SO}_4^{2-}$ $\text{PO}_4^{3-}$ Ammonium $\text{OH}^-$ $\text{CO}_3^{2-}$ carbonate $\text{CH}_3\text{COO}^-$ or $\text{C}_2\text{O}_2^-$

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