

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

losing

true

true

polyatomic

seven

true (released)

9. Calcium becomes a monatomic cation by gaining two electrons.

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

12. Monatomic ions consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

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Al Napier, Philip Judd



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Life Science, Grades 6-7 Tony Wright, 1994-07-13

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Table of Contents **7 1 Review And Reinforcement Answer Key**

1. Understanding the eBook **7 1 Review And Reinforcement Answer Key**
 - The Rise of Digital Reading **7 1 Review And Reinforcement Answer Key**
 - Advantages of eBooks Over Traditional Books
2. Identifying **7 1 Review And Reinforcement Answer Key**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **7 1 Review And Reinforcement Answer Key**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **7 1 Review And Reinforcement Answer Key**
 - Personalized Recommendations
 - **7 1 Review And Reinforcement Answer Key** User Reviews and Ratings
 - **7 1 Review And Reinforcement Answer Key** and Bestseller Lists

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 - 7 1 Review And Reinforcement Answer Key eBook Subscription Services
 - 7 1 Review And Reinforcement Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - 7 1 Review And Reinforcement Answer Key Compatibility with Devices
 - 7 1 Review And Reinforcement Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 7 1 Review And Reinforcement Answer Key
 - Highlighting and Note-Taking 7 1 Review And Reinforcement Answer Key
 - Interactive Elements 7 1 Review And Reinforcement Answer Key
8. Staying Engaged with 7 1 Review And Reinforcement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 7 1 Review And Reinforcement Answer Key
9. Balancing eBooks and Physical Books 7 1 Review And Reinforcement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 7 1 Review And Reinforcement Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 7 1 Review And Reinforcement Answer Key
 - Setting Reading Goals 7 1 Review And Reinforcement Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 7 1 Review And Reinforcement Answer Key
 - Fact-Checking eBook Content of 7 1 Review And Reinforcement Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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