

Steric # = Total e ⁻ pairs on central atom	Electron Distribution	# Bond Pairs	# Lone Pairs	General AXE Formula	Molecular Shape	e.g.	Lewis Structure	3-D Shape (Sketch)	Bond Angle(s)	Polar Molecule? P or NP (4.5)
2	Linear	1	0	AX	Linear	H ₂	H-H		180°	NP
4	Linear	2	0	AX ₂	Linear	BeCl ₂			180°	NP
3	Trigonal Planar	3	0	AX ₃	Trigonal Planar	BCl ₃			120°	NP
		2	1	AX ₂ E	"Bent"	SO ₂			<120°	P ≠ EN >0.4
4	Tetrahedral	4	0	AX ₄	Tetrahedral	CCl ₄			109.5°	NP
		3	1	AX ₃ E	Trigonal Pyramidal	NCl ₃			107°	P ≠ EN >0.4
		2	2	AX ₂ E ₂	Bent	OCl ₂ H ₂ O			104.5°	P ≠ EN >0.4

Dot Vsepr Lab Class Set

DP Hallahan



Dot Vsepr Lab Class Set:

Chemistry in the Laboratory James M. Postma, Julian L. Roberts, J. Leland Hollenberg, 2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry As always the new edition of Chemistry in the Laboratory offers precise easy to follow instructions helpful illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project

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