

Introduction

This course is a general chemistry course for students who are interested in pursuing a degree in a science field. The course covers the basic principles of chemistry, including atomic structure, chemical bonding, and chemical reactions. The course is designed to provide students with a solid foundation in chemistry, which is essential for understanding the physical and chemical world around us.

Prerequisites: Students must have completed high school chemistry or an equivalent course. Students who do not have the prerequisite should consult with their advisor for more information.

Part 1: Atomic Structure



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule.
2. Describe the structure of an atom, including the nucleus and the electron cloud. Explain the role of protons, neutrons, and electrons.
3. Explain the periodic table of elements, including the groups and periods. Describe the properties of the elements in each group.



4. Calculate the atomic mass of an element from its isotopes.
5. Calculate the molar mass of a compound from its chemical formula.

Element	Atomic Mass	Molar Mass	Group	Period	Block
H	1.008	1.008	1	1	s
He	4.003	4.003	18	1	p
Li	6.941	6.941	1	2	s
Be	9.012	9.012	2	2	p

6. Calculate the number of atoms in a given mass of an element.
7. Calculate the number of molecules in a given mass of a compound.

Part 2: Chemical Bonding



8. Explain the difference between ionic and covalent bonding. Describe the properties of ionic and covalent compounds.
9. Draw the Lewis structure of a molecule, showing the arrangement of atoms and the bonds between them.

$$2H_2 + O_2 \rightarrow 2H_2O$$

10. Calculate the percent composition of a compound from its chemical formula.

Compound	Element	Mass %	Formula	Mass %
H ₂ O	H	11.1%	H ₂ O	11.1%
H ₂ O	O	88.9%	H ₂ O	88.9%
CO ₂	C	27.3%	CO ₂	27.3%
CO ₂	O	72.7%	CO ₂	72.7%

Page 1 of 10 pages. Please refer to the back of the page for more information.

Stoichiometry Phet Lab Answers

N Colangelo



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