

Chapter Test B

Chapter: Stoichiometry

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. Knowing the mole ratio of a reactant and product in a chemical reaction would allow you to determine
- the energy released in the reaction.
 - the speed of the reaction.
 - the mass of the product produced from a known mass of reactant.
 - whether the reaction is reversible.
- _____ 2. In the equation $6\text{CO}_2 + 6\text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2$, the mole ratio of water to oxygen is
- 2:1.
 - 1:1.
 - 3:2.
 - 6:8.
- _____ 3. In the equation $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$, the mole ratio of calcium hydroxide to aluminum hydroxide is
- 1:3.
 - 2:3.
 - 1:1.
 - 3:2.
- _____ 4. For the reaction represented by the equation $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, how many moles of nitrogen are required to produce 18 mol of ammonia?
- 9
 - 18
 - 27
 - 36
- _____ 5. In the reaction represented by the equation $\text{A} + \text{B} \rightarrow \text{C} + \text{D}$, if there is more reactant B than is required to completely react with all of A, then
- A is the limiting reactant.
 - B is the limiting reactant.
 - there is no limiting reactant.
 - no product can be formed.

Stoichiometry Chapter Test B

Georgios Tsaparis



Stoichiometry Chapter Test B:

Chapter-wise Topical Objective Study Package for CBSE 2022 Class 12 Term I Chemistry Disha Experts, 2021-09-01 Disha s Chapter wise Topical Study Package for CBSE 2022 Class 12 Term I Chemistry is designed on the exact lines of the latest syllabus and paper pattern prescribed by the CBSE board Circular dated July 22 2021 for the Term I Exam to be held in November The Book consists of a total of 6 Chapters of Term I Each chapter is divided into 3 4 Topics Each Topic covers exhaustive theory with Illustrations followed by an Objective Exercise consisting of MCQs AR Case based VSA SA Questions Further the Chapter covers Concept Maps Important Formulae NCERT Exemplar Past Year Questions The Past Solved Objective Questions covered in the book are from 2021 CBSE Sample Paper 2020 2019 2018 2017 In the end of the Chapter an Objective Practice Exercise and a Chapter Test is provided for final practice and assessment There are a total of 1000 Objective Questions with Solutions The book is a One Stop Solution for Learning Practice Revision *Problems and Problem Solving in Chemistry Education* Georgios Tsaparlis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **43 JEE Main Chemistry Online (2019-2012) & Offline (2018-2002) Chapter-wise + Topic-wise Solved Papers 3rd Edition** Disha Experts, 2019-05-16 The book 43 JEE Main Chemistry Online Offline Topic wise Solved Papers provides the last 18 years ONLINE OFFLINE 2002 18 papers The book contains a total of 43 papers 17 papers of JEE Main from the year 2002 2018 held OFFLINE including the AIEEE 2011 RESCHEDULED paper and 25 JEE Main papers held ONLINE from 2012 19 The book also provides separate web link free access to the 16 Online Solved Papers held in January April 2019 The book is distributed into around 30 Chapters exactly following the chapter sequence of the NCERT books of class 11 and 12 The questions in each Chapter are further divided into 2 3 topics The Questions are immediately followed by their detailed solutions The book constitutes of 1680 MCQs with Solutions **Arun Deep's**

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subsequent ΔH values $2 \text{ CO}_2 \text{ (g)}$, $2 \text{ O}_2 \text{ (g)}$, $2 \text{ C}_2\text{H}_2 \text{ (g)}$, $5 \text{ O}_2 \text{ (g)}$, $2 \text{ H}_2 \text{ (g)}$, $2 \text{ C}_2\text{H}_6 \text{ (g)}$, $\Delta H = 94.5 \text{ kJ}$, $\text{H}_2\text{O (g)}$, $\text{H}_2 \text{ (g)}$, $\frac{1}{2} \text{ O}_2 \text{ (g)}$, $\Delta H = 71.2 \text{ kJ}$, $\text{C}_2\text{H}_6 \text{ (g)}$, $7 \text{ O}_2 \text{ (g)}$, $2 \text{ CO}_2 \text{ (g)}$, $3 \text{ H}_2 \text{ (g)}$, $\text{O}_2 \text{ (g)}$, $\Delta H = 283 \text{ kJ}$, $\text{C}_2\text{H}_6 \text{ (g)}$

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enthalpy for naoh s hess s law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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web thus for reaction 9 2 in which 1 mole of h_3o^+ and 1 mole of oh^- combine to form 2 moles of h_2o $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{products} - \Delta h_{reactants}$ is positive if heat is absorbed that is if $\Delta h_{products}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$ 2 $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$ 3