

Biology Exam #2 (Chapters 2, 3, 4, 6, 7, 8)

Chapter 2:

Atoms are made up of electrons (- outside the nucleus), protons(+inside the nucleus), and neutrons (0 inside the nucleus)

Atomic number= number of protons

Mass number=protons + nucleus

Matter is composed of elements and exists as solid, liquid, gas, or plasma

Element: a substance that cannot be broken down by chemical means to a simpler substance that has different physical properties.

Basic elements to life include C, H, N, O, P, S

Life's diversity is based on carbon, almost all the molecules of a cell are composed of carbon

Organic molecules: carbon-based molecules, molecules have C and H, those with only C and H are called hydrocarbons

Isotopes: atoms of an element with different number of neutrons

Radioactive isotope emits energy as rays and subatomic particles

Uses of radiation include medical imaging, killing bacteria and fungi, killing cancer cells

For neutral atoms, the number of valence electrons is equal to the atom's main group number

ATP: currency of the cell, high energy molecule, undergoes hydrolysis which releases energy

Polar covalent bonds: sharing between two atoms that is unequal

Oxygen is more electronegative than Hydrogen

Second strongest

Nonpolar covalent bonds: sharing of electrons between two atoms is fairly equal

Strongest

Carbon bonds: four electrons on outer shell

Electronegativity: attraction of an atom for electrons in covalent bond

Hydrogen bonding: attraction between negative oxygen and positive hydrogen

Bonds is weak individually, but strong together(weakest)

Biology Chapter 34 Study Guide Answers

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Table of Contents Biology Chapter 34 Study Guide Answers

1. Understanding the eBook Biology Chapter 34 Study Guide Answers
 - The Rise of Digital Reading Biology Chapter 34 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Biology Chapter 34 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Biology Chapter 34 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Biology Chapter 34 Study Guide Answers
 - Personalized Recommendations
 - Biology Chapter 34 Study Guide Answers User Reviews and Ratings
 - Biology Chapter 34 Study Guide Answers and Bestseller Lists

5. Accessing Biology Chapter 34 Study Guide Answers Free and Paid eBooks
 - Biology Chapter 34 Study Guide Answers Public Domain eBooks
 - Biology Chapter 34 Study Guide Answers eBook Subscription Services
 - Biology Chapter 34 Study Guide Answers Budget-Friendly Options
6. Navigating Biology Chapter 34 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Biology Chapter 34 Study Guide Answers Compatibility with Devices
 - Biology Chapter 34 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Biology Chapter 34 Study Guide Answers
 - Highlighting and Note-Taking Biology Chapter 34 Study Guide Answers
 - Interactive Elements Biology Chapter 34 Study Guide Answers
8. Staying Engaged with Biology Chapter 34 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Biology Chapter 34 Study Guide Answers
9. Balancing eBooks and Physical Books Biology Chapter 34 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Biology Chapter 34 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Biology Chapter 34 Study Guide Answers
 - Setting Reading Goals Biology Chapter 34 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Biology Chapter 34 Study Guide Answers
 - Fact-Checking eBook Content of Biology Chapter 34 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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