

Advanced Chemistry useful handout #1

Metric conversion chart

10^9 g =	1 Gg	giga	
10^6 g =	1 Mg	mega	
10^3 g =	1 kg	kilo	$1000 \text{ g} = 1 \text{ kg}$
1 g =	1 g	BASE	
1 g =	10^1 dg	deci	$1 \text{ g} = 10 \text{ dg}$
1 g =	10^2 cg	centi	$1 \text{ g} = 100 \text{ cg}$
1 g =	10^3 mg	milli	$1 \text{ g} = 1000 \text{ mg}$
1 g =	10^6 μ g	micro	
1 g =	10^9 ng	nano	
1 g =	10^{12} pg	pico	

10^{10} Å (Angstroms) = 1 meter (atomic size)

Linear to Volume: When converting between linear and volume measurements, you need to *cube* the conversion factor! $10 \text{ cm} = 1 \text{ dm}$; so $(10 \text{ cm})^3 = (1 \text{ dm})^3$ thus $1000 \text{ cm}^3 = 1 \text{ dm}^3 = 1 \text{ Liter}$

Some Helpful Formulas:

Gas Laws: Conversions: STP for gases = 0°C and 1 atm ; (or 273 K , 760 mm Hg)

Combined : $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$ Diffusion: $\frac{\text{rate A}}{\text{rate B}} = \frac{\sqrt{FW B}}{\sqrt{FW A}}$ Molar volume of = $\frac{22.4 \text{ L}}{\text{mole}}$
 a gas at STP

Ideal Gas Law: $PV = nRT$ If you know P, V, T of a gas, you can calculate moles!

For $PV = nRT$: P in atm, V in L, n = moles, T in K, Ideal Gas Constant, R = $0.0821 \text{ L atm/mol K}$
 (use Partial Pressure of a gas in $PV = nRT$ to get mol of that particular gas)

Dalton's law of Partial Pressures: $P_{\text{total}} = P_a + P_b + P_c + \dots$ $P_{\text{dry gas}} = P_{\text{total}} - P_{\text{water vapor}}$

Partial pressure = $P_a = X_a P_{\text{total}}$ where mol fraction = $X_a = \frac{\text{mol a}}{\text{Total mol of gases}}$

Solutions: Molarity = $\frac{\text{moles solute}}{\text{L of solution}}$ or $M = \frac{\text{mol}}{\text{L}}$ mol = V M $V = \frac{\text{mol}}{M}$ (V must be in Liters)

Or : can use millimoles and milliliters: $M = \frac{\text{mmol}}{\text{mL}}$ mmol = mL M $\text{mL} = \frac{\text{mmol}}{M}$

Dilutions: $V_1 M_1 = V_2 M_2$ (can use mL for Volumes) Acid Base Titration: $V_a M_a (\#H) = V_b M_b (\#OH)$

molality = $m = \frac{\text{mol solute}}{\text{kg solvent}}$ mol fraction = $X_a = \frac{\text{mol a}}{\text{total mol}}$ Weight % = $\frac{\text{g solute}}{100 \text{ g solution}}$

Additional conversion factors and constants:

Energy:

$4.18 \text{ joules} = 1 \text{ calorie}$ $1 \text{ Joule} = \frac{1 \text{ kg m}^2}{\text{sec}^2}$

Volume: $1 \text{ cm}^3 = 1 \text{ mL}$ $1 \text{ dm}^3 = 1 \text{ L}$

$1 \text{ m}^3 = 10^3 \text{ L}$

Density = $\frac{\text{Mass}}{\text{Volume}}$; $\frac{1 \text{ gram}}{\text{mL}}$ (water only)

Avogadro's number: $N = \frac{6.02 \times 10^{23} \text{ particles}}{1 \text{ mole}}$

Kelvin = $^\circ\text{Celsius} + 273$

$1 \text{ atm} = 760 \text{ mm Hg}$

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