

## Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ , with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The  $K_a$  for butanoic acid is  $1.54 \times 10^{-5}$ .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M  $\text{HClO}_4$  after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M  $\text{Ba}(\text{OH})_2$  solution have been added to 25.00 mL of 0.0400 M  $\text{HNO}_3$  solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

# Acid Base Titration Problems With Answers

**Ying Liu**



## **Acid Base Titration Problems With Answers:**

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**Programmed Problem-solving for First-year Chemistry** Arnold B. Loebel, 1983

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*Applied Physical Pharmacy, Third Edition* Mansoor Amiji, Thomas J. Cook, W. Cary Mobley, 2019-06-26 A complete practice oriented introduction to physical pharmacy Written to clearly and simply explain how drugs work this textbook explores the fundamental physicochemical attributes and processes important for understanding how a drug is transformed into a usable product that is administered to a patient to reach its pharmacological target and then exists the body Applied Physical Pharmacy Third Edition begins with a review of the key biopharmaceutics concepts of drug liberation absorption distribution metabolism and excretion These concepts and others set the framework for the subsequent chapters that describe physicochemical properties and process related to the fate of the drug Other physical pharmacy topics important to drug formulation are discussed in the chapters that follow which describe dispersal systems interfacial phenomena and rheology The textbook concludes with an overview of the principles of kinetics that are important for understanding the rates at which many of the processes discussed in previous chapters occur Chapters in this Third Edition retain the acclaimed learning aids of previous editions including Learning Objectives Practice Problems Key Points and Clinical Questions In order to be of greater value to the pharmacy student more clinical questions have been added and many tables have been updated with more current products and excipients Understanding Chemistry Fred M. Dewey, 1994 *Principles of Modern Chemistry*

David W. Oxtoby, H. P. Gillis, Norman H. Nachtrieb, 2002 This successful textbook continues to address students in honours and high mainstream general chemistry courses The chapters in the fourth edition are organised within larger units to emphasise the conceptual structure of chemistry

**Chemistry Problems and how to Solve Them** Paul Reheard Frey, 1969 Calculations in Analytical Chemistry Quintus Fernando, Michael D. Ryan, 1982 **A Learning Program for Chemistry** Peter Demmin, 1993-11 **Chemical Problem-solving by Dimensional Analysis** Arnold B. Loebel, 1978 5 Steps to a 5: AP Chemistry 2022 Elite Student Edition Mary Millhollon, Richard H. Langley, 2021-08-04 MATCHES THE LATEST EXAM Let us supplement your AP classroom experience with this multi platform study guide The immensely popular 5 Steps to a 5 AP Chemistry Elite Student Edition has been updated for the 2021 22 school year and now contains 3 full length practice exams available in the book and online that reflect the latest exam 5 Minutes to a 5 section with a 5 minute

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