



Accounting Principles 9th Edition Solutions

Mark Steyvers



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web the six no₃⁻ ions and the six na⁺ ions that appear on both sides of the equation are spectator ions that can be

canceled to give the net ionic equation $3\text{Ba}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Ba}_3(\text{PO}_4)_2(\text{s})$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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web 18 terms lourdes l22 preview terms in this set 4 molecular equation showing the complete chemical formulas of reactions and products $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$ complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no₃ aq ions stay as na aq ions and no₃ aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be 2cl aq cu₂ aq cucl₂ s 2cl aq cu₂ aq cucl₂ s these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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