Chemical Kinetics

4.1 Rate of a Chemical Reaction

1. For the chemical reaction,
   $N_{2(g)} + 3H_{2(g)} \rightleftharpoons 2NH_{3(g)}$
   the correct option is

   (a) $3\dfrac{d[H_2]}{dt} = 2\dfrac{d[NH_3]}{dt}$

   (b) $-\dfrac{1}{3}\dfrac{d[H_2]}{dt} = -\dfrac{1}{2}\dfrac{d[NH_3]}{dt}$

   (c) $-\dfrac{d[N_2]}{dt} = 2\dfrac{d[NH_3]}{dt}$

   (d) $-\dfrac{d[N_2]}{dt} = \dfrac{1}{2}\dfrac{d[NH_3]}{dt}$ (NEET 2019)

2. The rate of the reaction : $2N_2O_5 \longrightarrow 4NO_2 + O_2$ can be written in three ways.

   $\dfrac{-d[N_2O_5]}{dt} = k[N_2O_5]$

   $\dfrac{d[NO_2]}{dt} = k'[N_2O_5]; \dfrac{d[O_2]}{dt} = k''[N_2O_5]$

   The relationship between k and k' and between k and k'' are

   (a) $k' = 2k, k'' = k$ (b) $k' = 2k, k'' = k/2$
   (c) $k' = 2k, k'' = 2k$ (d) $k' = k, k'' = k$

   (Mains 2011)

3. For the reaction $N_2O_{5(g)} \longrightarrow 2NO_{2(g)} + 1/2O_{2(g)}$ the value of rate of disappearance of N_2O_5 is given as 6.25×10^{-3} mol L^{-1}s^{-1}. The rate of formation of NO_2 and O_2 is given respectively as :
   (a) 6.25×10^{-3} mol L^{-1} s^{-1} and
       6.25×10^{-3} mol L^{-1} s^{-1}
   (b) 1.25×10^{-2} mol L^{-1} s^{-1} and
       3.125×10^{-3} mol L^{-1} s^{-1}
   (c) 6.25×10^{-3} mol L^{-1} s^{-1} and
       3.125×10^{-3} mol L^{-1} s^{-1}
   (d) 1.25×10^{-2} mol L^{-1} s^{-1} and
       6.25×10^{-3} mol L^{-1} s^{-1}

   (2010)

4. For the reaction, $N_2 + 3H_2 \longrightarrow 2NH_3$, if $\dfrac{d[NH_3]}{dt} = 2 \times 10^{-4}$ mol L^{-1} s^{-1}, the value of $-\dfrac{d[H_2]}{dt}$ would be

   (a) 4×10^{-4} mol L^{-1} s^{-1} (b) 6×10^{-4} mol L^{-1} s^{-1}
   (c) 1×10^{-4} mol L^{-1} s^{-1} (d) 3×10^{-4} mol L^{-1} s^{-1}

   (2009)

5. In the reaction,
   $BrO_3^-{}_{(aq)} + 5Br^-{}_{(aq)} + 6H^+ \longrightarrow 3Br_{2(l)} + 3H_2O$
   The rate of appearance of bromine (Br_2) is related to rate of disappearance of bromide ions as

   (a) $\dfrac{d[Br_2]}{dt} = -\dfrac{5}{3}\dfrac{d[Br^-]}{dt}$ (b) $\dfrac{d[Br_2]}{dt} = \dfrac{5}{3}\dfrac{d[Br^-]}{dt}$

   (c) $\dfrac{d[Br_2]}{dt} = \dfrac{3}{5}\dfrac{d[Br^-]}{dt}$ (d) $\dfrac{d[Br_2]}{dt} = -\dfrac{3}{5}\dfrac{d[Br^-]}{dt}$

   (2009)

6. Consider the reaction :
   $N_{2(g)} + 3H_{2(g)} \longrightarrow 2NH_{3(g)}$
   The equality relationship between $\dfrac{d[NH_3]}{dt}$ and $-\dfrac{d[H_2]}{dt}$ is

   (a) $\dfrac{d[NH_3]}{dt} = -\dfrac{d[H_2]}{dt}$

   (b) $\dfrac{[NH_3]}{dt} = -\dfrac{1}{3}\dfrac{d[H_2]}{dt}$

   (c) $+\dfrac{d[NH_3]}{dt} = -\dfrac{2}{3}\dfrac{d[H_2]}{dt}$

   (d) $+\dfrac{d[NH_3]}{dt} = -\dfrac{3}{2}\dfrac{d[H_2]}{dt}$

   (2006)

7. For the reaction, $2A + B \longrightarrow 3C + D$, which of the following does not express the reaction rate?

   (a) $-\dfrac{d[A]}{2dt}$ (b) $-\dfrac{d[C]}{3dt}$

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