

# AS Level Chemistry

## 11: Group 17

ALLEN HARRING  
CAPITAL

- $I_2$  (s) : pale yellow gas  
 $Cl_2$  (g) : pale green gas  
 $Br_2$  (l) : volatile brown liquid  
 $I_2$  (s) : black solid / purple vapour
- } colour intensity increases down Grp 17.

- 6 p block elements  
 oxidation number: -1 or 0. (can go up to +7 bc of vacant d orbitals - except F)

### Atomic Radius

Trends change overwhelmed by 7 shielding, nuclear attraction remains same  $Z$  number, **atomic radius ↑**

### MP/BP & Volatility

simple molecular structures → weak D-D than molecules → no. of electrons ↑ → **boil pt ↑** → **mp/bp ↑**  
 → diff. to separate molecules → **volatility ↓**

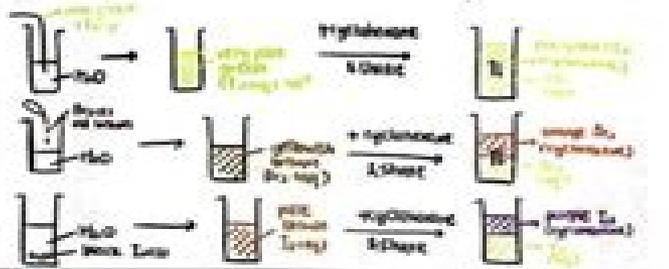
### Solubility in water

oxidising power ↓ → solubility ↓

- F: - oxides water vigorously even at -100°C.  
 $2F_2(g) + 2H_2O(l) \rightarrow 4HF(g) + O_2(g)$   
 Cl: - disproportionates partially in water  
 $Cl_2(g) + H_2O(l) \rightleftharpoons HCl(aq) + HClO(aq)$   
 - chlorides are decomposed w/ strong oxidant  
 $2HClO(aq) \rightarrow 2HCl(aq) + O_2(g)$   
 Br: - moderately soluble in water → forms equilibrium with mixture of  $Br_2(aq)$  -  $H_2O$  formed negligible  
 I: - dissolves slightly in water to give pale yellow soln.  
 - soluble in  $KI$  sol<sup>n</sup> to form  $I_3^-$  (brown sol<sup>n</sup>)  
 $I_2(s) + I^-(aq) \rightarrow I_3^-(aq)$

### Solubility in Cyclohexane

- insoluble in water → forms 2 separate layers  
 - halogens dissolved → distinctly diff. colours



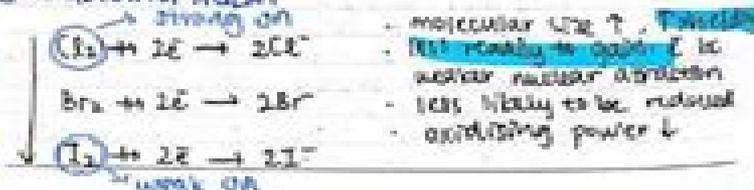
### Bond Energy

- atomic radii ↑ → **less effective orbital overlap** → less attraction for shared pair → increasing bond length → **bond energy ↓** from Cl to I.

Why F-F bond isn't longest:

- smallest valence orbitals → LP of  $F$  repel each other → weakens F-F bond (like phosphorus w/ arsenic soln increases bc non bonding electrons have more space to spread out)

### 11.2 OXIDISING AGENT



straight **A\*** student's notes

# A Level Chemistry Revision Notes

**AW Chickering**



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