

Prentice Hall Physical Science Chapter 4 Atomic Structure

4.1 Studying Atoms

- A. **Ancient Greek Models of Atoms**
 - Democritus believed all matter was made of extremely small particles that could not be divided
 - he called the particles atoms (from the Greek atomos meaning indivisible)
 - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. **Dalton's Atomic Theory**
 - he proved atoms exist by discovering that compounds have a fixed composition
 - Dalton's Theory:
 - all elements are composed of atoms
 - all atoms of the same element have the same mass, and atoms of different elements have different masses
 - compounds contain atoms of more than one element
 - in a particular compound, atoms of different element always combine in the same way
 - he pictured atoms as solid spheres
 - in time it was found that not all of this was completely correct
- C. **Thomson's Model of the Atom**
 - some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
 - Thomson used an electric current to learn more about atoms
 - he passed an electric current through a gas and a glowing beam appeared
 - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
 - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
 - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
 - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. **Rutherford's Atomic Theory**
 - Rutherford discovered positively charged, fast moving particles called alpha particles
 - he shot some of these particles a piece of gold foil, thinking they would go straight through
 - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
 - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
 - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
 - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

Chapter Assessment Prentice Hall Physical Science

TD Snyder



Chapter Assessment Prentice Hall Physical Science:

The book delves into Chapter Assessment Prentice Hall Physical Science. Chapter Assessment Prentice Hall Physical Science is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chapter Assessment Prentice Hall Physical Science, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chapter Assessment Prentice Hall Physical Science
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 - Chapter 4: Chapter Assessment Prentice Hall Physical Science in Specific Contexts
 - Chapter 5: Conclusion
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