

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are ${}^{32}_{16}\text{S}$ and ${}^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 41
2. sub atomic 42
3. negative 42
4. protons 42
5. neutrons 42
6. nucleus 43
7. Rutherford 43
8. positive 43
9. electrons 43
10. protons (elec) 44
11. number 44
12. electrons 44
13. protons (elec) 45
14. neutrons 45
15. isotopes 45
16. atomic mass 47
17. 16 45
18. 16 46
19. 16 46
20. 18 45

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C Cleary



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Table of Contents Chapter 4 Atomic Structure Answer Key

1. Understanding the eBook Chapter 4 Atomic Structure Answer Key
 - The Rise of Digital Reading Chapter 4 Atomic Structure Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Chapter 4 Atomic Structure Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Answer Key
 - Personalized Recommendations
 - Chapter 4 Atomic Structure Answer Key User Reviews and Ratings
 - Chapter 4 Atomic Structure Answer Key and Bestseller Lists

5. Accessing Chapter 4 Atomic Structure Answer Key Free and Paid eBooks
 - Chapter 4 Atomic Structure Answer Key Public Domain eBooks
 - Chapter 4 Atomic Structure Answer Key eBook Subscription Services
 - Chapter 4 Atomic Structure Answer Key Budget-Friendly Options
6. Navigating Chapter 4 Atomic Structure Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Answer Key Compatibility with Devices
 - Chapter 4 Atomic Structure Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Answer Key
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Answer Key
 - Interactive Elements Chapter 4 Atomic Structure Answer Key
8. Staying Engaged with Chapter 4 Atomic Structure Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Answer Key
9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Atomic Structure Answer Key
 - Setting Reading Goals Chapter 4 Atomic Structure Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Atomic Structure Answer Key
 - Fact-Checking eBook Content of Chapter 4 Atomic Structure Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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