

5.6 Writing Formulas for Ionic Compounds Assignment

Instructions:

For this assignment, you are being asked to write formulas for Ionic Compounds and for Ionic Compounds containing Polyatomic Ions. Review the examples for each section and then write the formula for the identified compounds.

Writing Formulas for Ionic Compounds Assignment

Examples:

Write the formula of potassium bromide.

- K^{+1} Br^{-1}
- The sum of the charges is $+1 + (-1) = 0$, therefore, the formula is KBr

Write the formula of calcium bromide.

- Ca^{+2} Br^{-1}
- The sum of the oxidation numbers is $+2 + (-1) = +1$, therefore, we need a second bromide (-1) to obtain a proper formula.
- The new sum is $+2 + 2(-1) = 0$, therefore, the formula is $CaBr_2$.

A simpler way to understand how to write formulas is to use the crisscross & reduce method. If there is only one atom of an element in the compound, you do not need to put the number 1 behind it. The following are examples:



Questions 1-6.

Write formulas for the following compounds:

1. sodium and fluorine



4. aluminum and sulfur



2. magnesium and iodine



5. hydrogen and oxygen



3. barium and oxygen



6. lithium and nitrogen



Assignment 5 Ionic Compounds

Xiaolong Qi



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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations *The Analytical Teacher* Victor P. Maiorana, 1994-07 *Using Physical Models of Biomolecules to Teach Concepts of Biochemical Structure in Introductory Undergraduate Chemistry* John Yi, 2004 **Journal of the American Chemical Society** American Chemical Society, 1960 Proceedings of the Society are included in v 1 59 1879 1937 **Antimony: Compounds of pentavalent antimony with six, five, and four Sb-C bonds** Markus Wieber, 1982 Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface **Canadian Journal of Chemistry**, 1961 *Virtual Chemlab* Brian F. Woodfield, Matthew C. Asplund, Steven Haderlie, 2006 Contains a full virtual lab environment as well as the pre arranged labs that are refer enced in the workbook and at the end of the chapter in the textbook Virtual ChemLab can be run directly from the CD or installed on the student s computer *Foundations of College Chemistry* Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Science Abstracts**, 1965 *Holt Chemistry* Ralph Thomas Myers, 2004 The Elements of Chemistry Lawrence Powell Eblin, 1965 Chemical concepts and the atomic theory Stoichiometry and the mole concept Atomic numbers and the periodic law Electronic configurations of the elements Chemical bonding and oxidation states Gas laws and the kinetic theory The liquid and solid states Oxygen Hydrogen and water Acids and bases Solutions and colloids Reaction rates and equilibrium Oxidation reduction and electrochemistry Energy and chemical change Organic compounds Biochemical compounds Transmutation of the elements N Nitrogen Reinhard Haubold, Peter Merlet, Claudia Heinrich-Sterzel, Astrid Wietelmann, Ulrike Ohms-Bredeman, Carol Strametz, 1993-10-20 This volume deals with binary nitrogen hydrogen compounds having two three or more nitrogen atoms with the exception of hydrazine and with compounds composed of nitrogen

hydrogen and noble gases The important species containing two nitrogen atoms N_2H N_2H_2 and N_2H_3 are described in the first part of this volume Next chains and cycles consisting of three nitrogen atoms are covered Among them hydrogen azide or hydrozoic acid HN_3 is the most extensively studied nitrogen hydrogen compound described in this volume With increasing number of nitrogen atoms the thermochemical stability declines There is however a considerable amount of information on molecules with up to nine linked nitrogen atoms Several of these binary nitrogen hydrogen compounds could only be isolated in the form of organic derivatives In that case data available for the organic derivatives were included if they were characteristic for the particular unsubstituted N H parent compound Comptes rendus de l'Académie bulgare des sciences, 1976 **Australian Journal of Chemistry**, 1985 **Russian Journal of Inorganic Chemistry**, 1990

Abstracts of Papers American Chemical Society. Meeting, **Concepts And Problems In Physical Chemistry** P.S. Raghavan, 1997 Contents Introduction Atoms Molecules and Formulas Chemical Equations and Stoichiometry Aqueous Reactions and Solution Stoichiometry Gases Intermolecular Forces Liquids and Solids Atoms Structure and the Periodic Table Chemical Bonding Chemical Thermodynamics Solutions Chemical Kinetics Chemical Equilibrium Acids and Bases Ionic Equilibria I Ionic Equilibria II Redox Reactions Electrochemistry Nuclear Chemistry *General Chemistry* Ralph Sherman Becker, Wayne Ernest Wentworth, 1980 *Quarterly Journal of the Chemical Society of London* Chemical Society (Great Britain), 1971

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